



KAPITAŁ LUDZKI
NARODOWA STRATEGIA SPÓJNOŚCI

UNIA EUROPEJSKA
EUROPEJSKI
FUNDUSZ SPOŁECZNY



BIOPHYSICS

**Prezentacja multimedialna współfinansowana przez
Unię Europejską w ramach
Europejskiego Funduszu Społecznego w projekcie pt.
*„Innowacyjna dydaktyka bez ograniczeń - zintegrowany
rozwój Politechniki Łódzkiej - zarządzanie Uczelnią,
nowoczesna oferta edukacyjna i wzmacniania zdolności
do zatrudniania osób niepełnosprawnych”***



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Lecture 8

Properties of the matter (8)

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Electrical Properties

Electrical properties depend on:

- electrical conductivity $\sigma = 1/\rho$ where ρ is a resistivity, the conductivity depends on:
 - type and concentration of electric charge
 - mobility of the charge in the electric field
- dielectric constant (relative permittivity) ϵ , the dielectric constant depends on:
 - space configuration of electric charge
 - ability of the charge to move



Electrical Properties

Gases - due to a lack of free charges and a weak molecular interactions at normal conditions generally are considered as not conductive materials. But under a high voltage, or temperature, gases can be converted into plasma and then are very good conductors (example: plasma lamps)

Liquids - due to a weak cohesion forces are also considered as not conductive materials. But some liquids are conductors of the second type. For example water is able to dissolve several salts, acids or bases, and can dissociate them to ions. Then we have electrolytes which are conductors. Chemically pure water is not conductive liquid - it is a dielectric substance.



Solid materials - can be found as conductors, semiconductors or insulators (dielectrics).

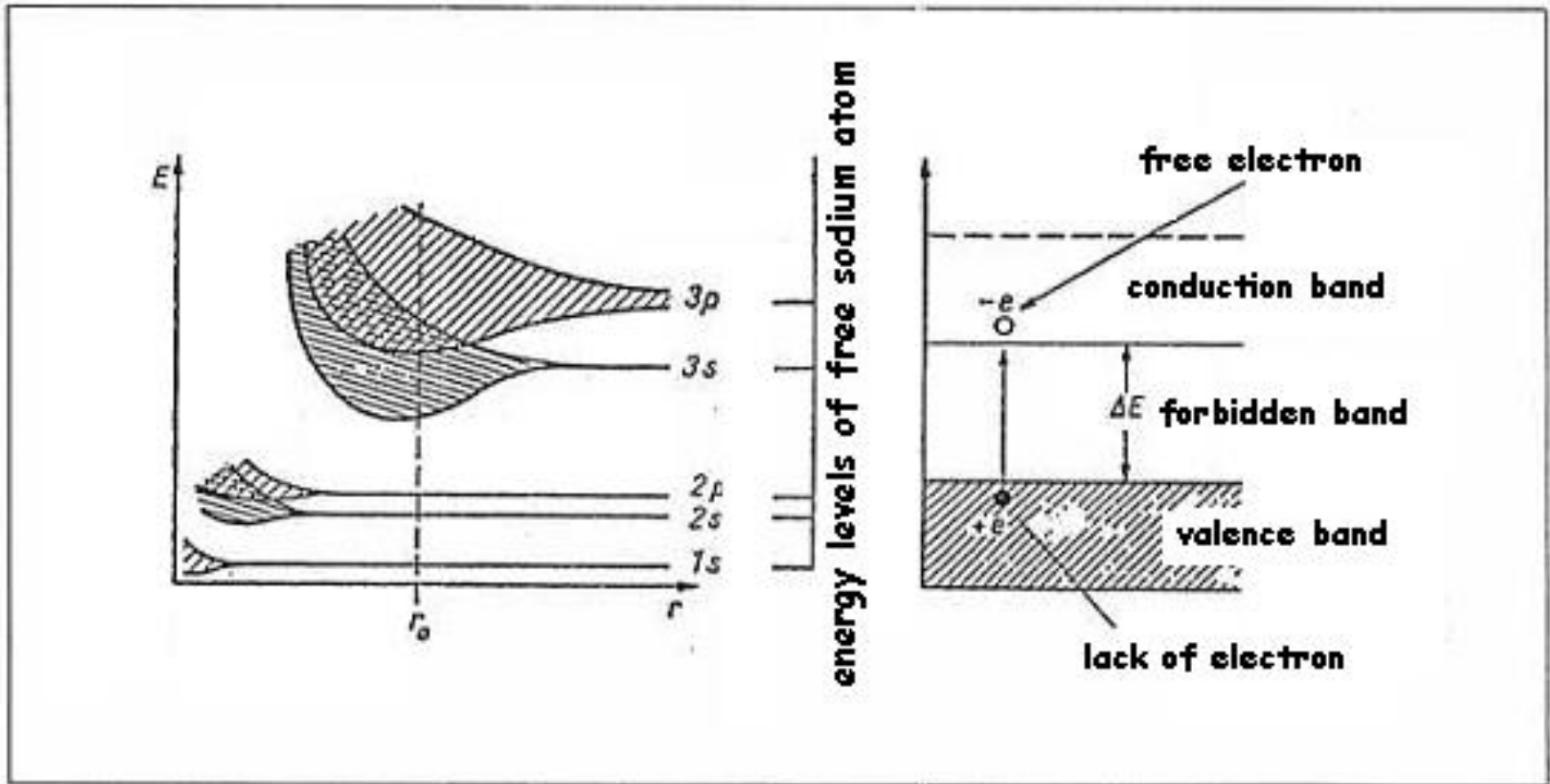
The property depends on the electronic structure of the substance, i.e. on the presence or absence, and on configuration of bandgaps.

For conductors no bandgaps between the valence band and conduction band is observed. For semiconductors this bandgap is rather small, below 2eV , whereas for insulators the bandgap is wider than 2eV .

This phenomenon is explained by the band theory of solids.



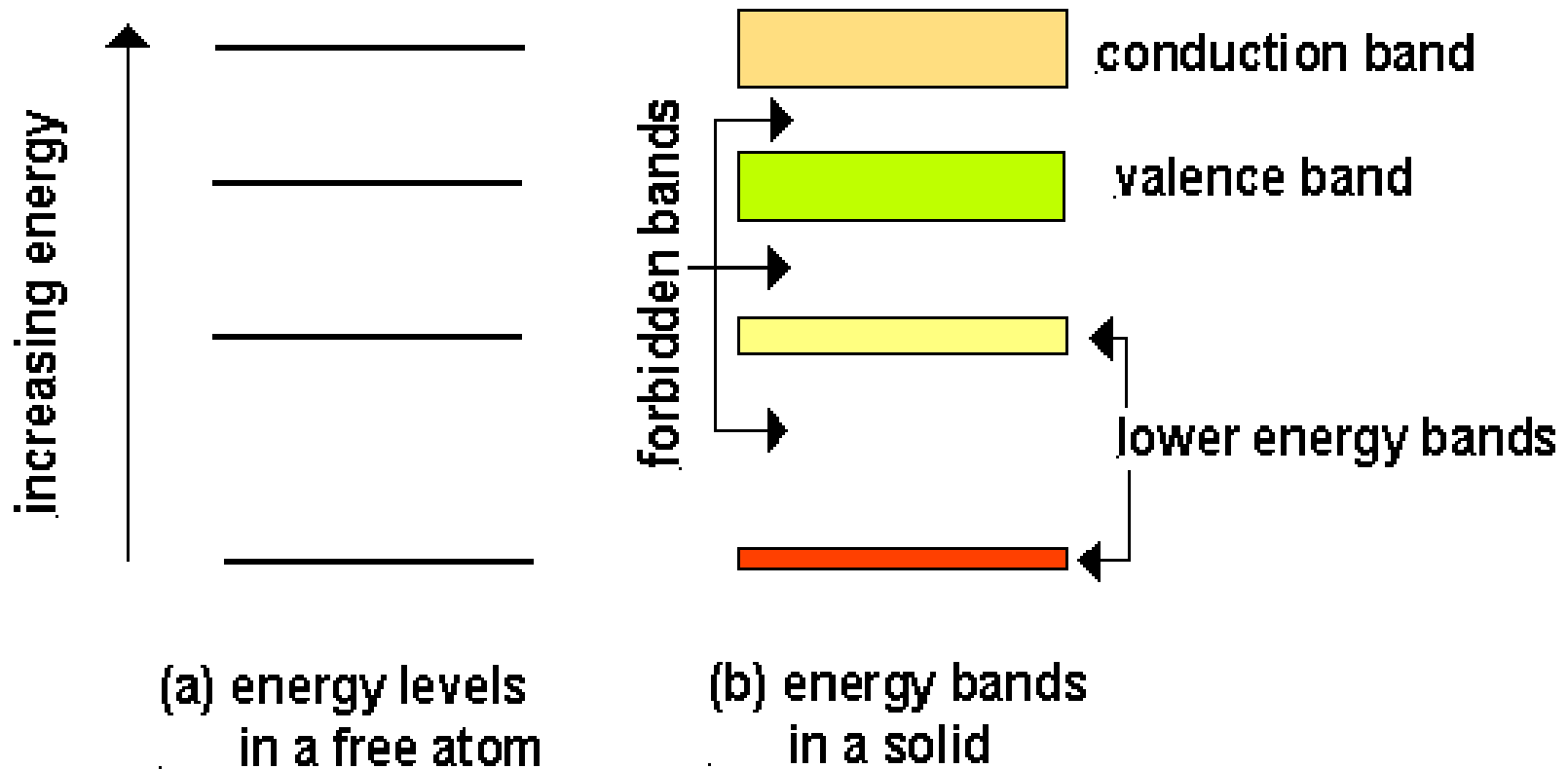
The Band Theory of Solids



Source: A. Piławski Podstawy Biofizyki

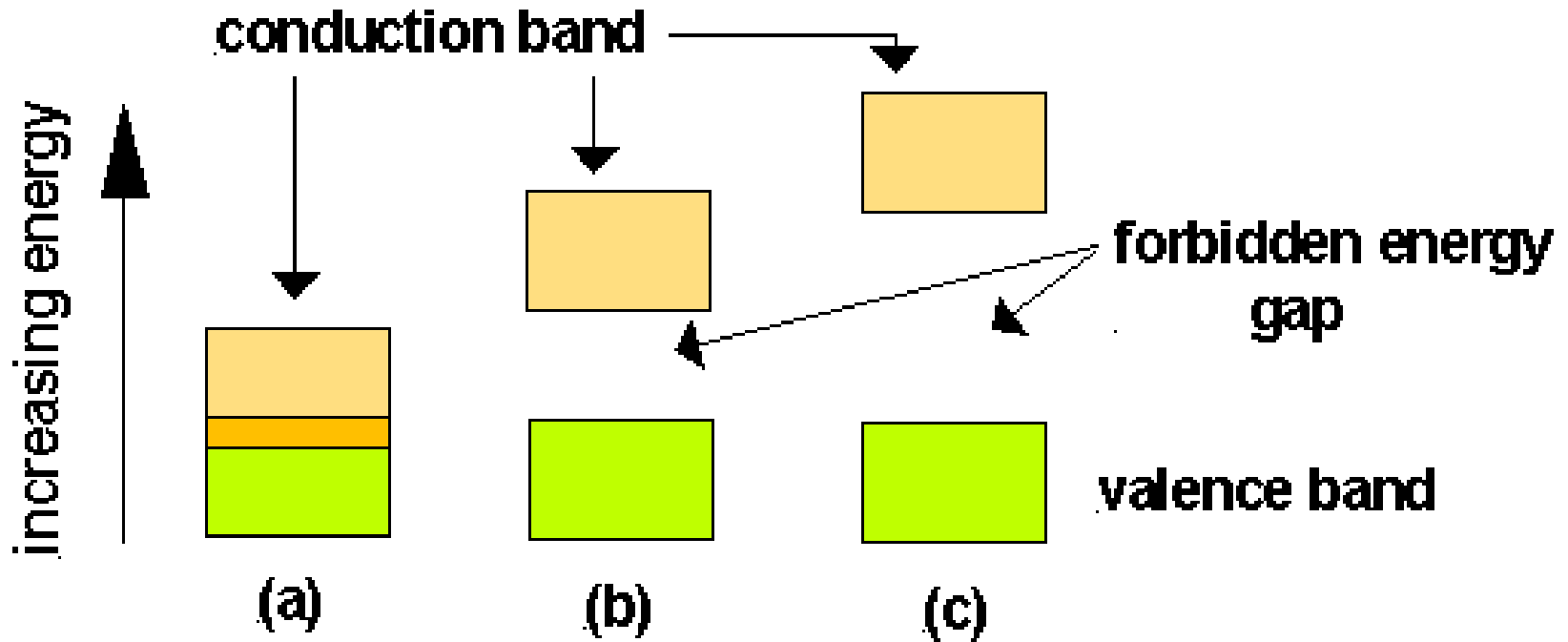


The Band Theory of Solids



Source: INTERNET

The Band Theory of Solids

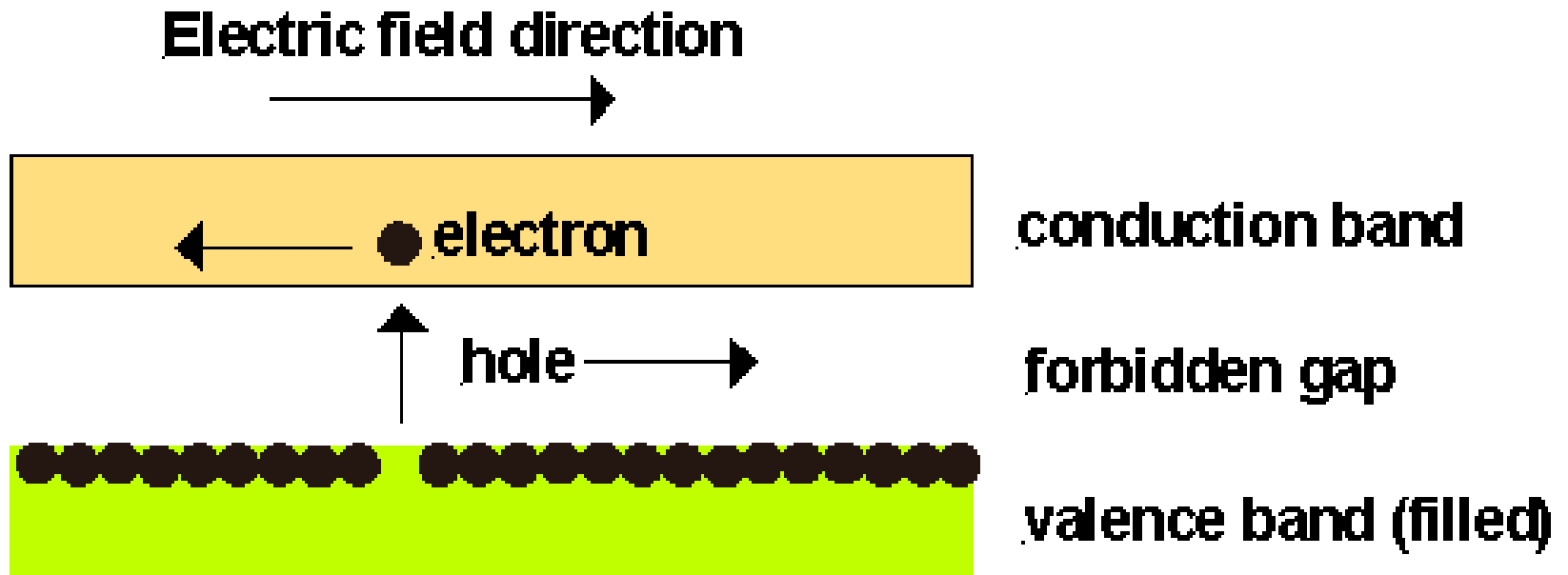


Energy bands in (a) conductors, (b) semiconductors & (c) insulators

Source: INTERNET



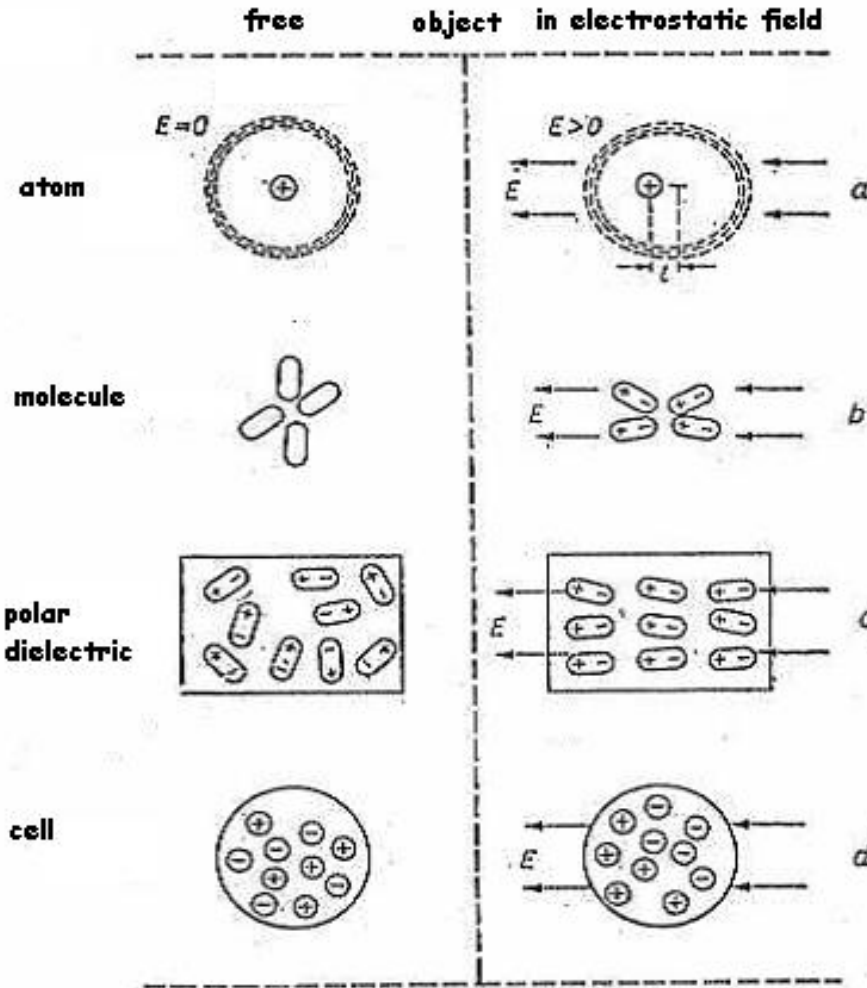
Semiconductors



Conduction via electrons & holes in a semiconductor

Source: INTERNET

Polarizations of Dielectrics



deforming electron
polarization

deforming atom
polarization

oriented
polarization

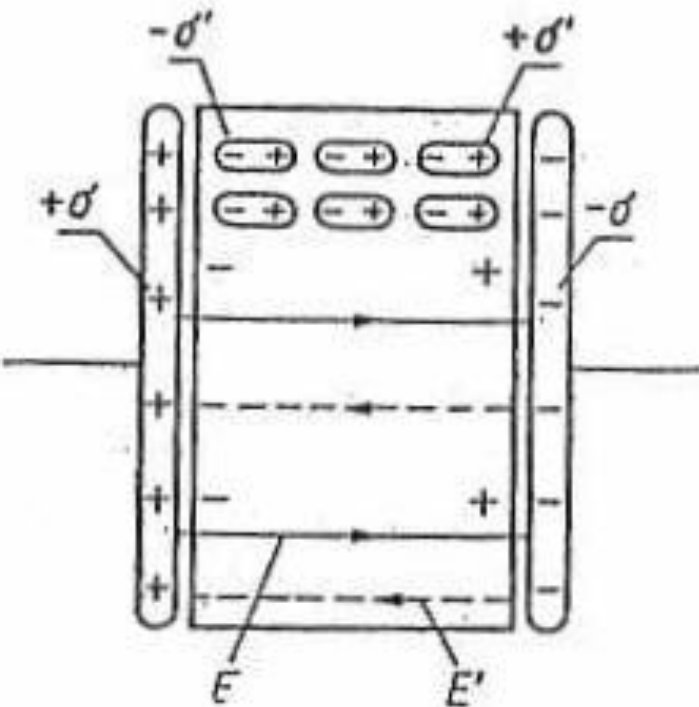
ionic
polarization

Source: A. Piławski Podstawy Biofizyki

Dielectrics in an Electric Field

Only an ideal dielectric has infinite resistance for DC current.

A real dielectric always presents a residual conductivity due to a residual impurity.



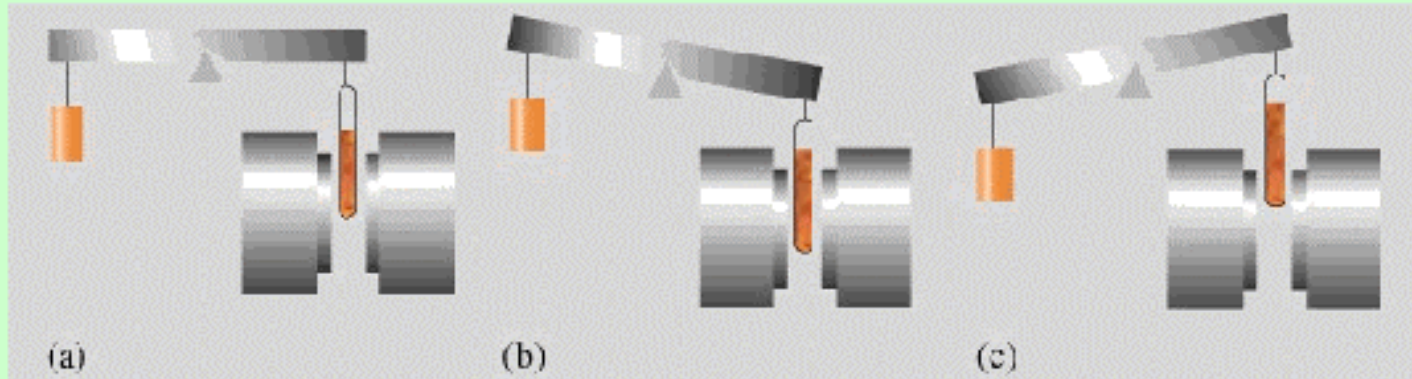
In a high frequency electric field dielectric dipoles can not keep up field changes due to friction, but can oscillate. This oscillations results in an increase in temperature.

This is why we can use microwave stove.

Source: A. Piławski Podstawy Biofizyki

Magnetic Properties

Paramagnetism and Diamagnetism

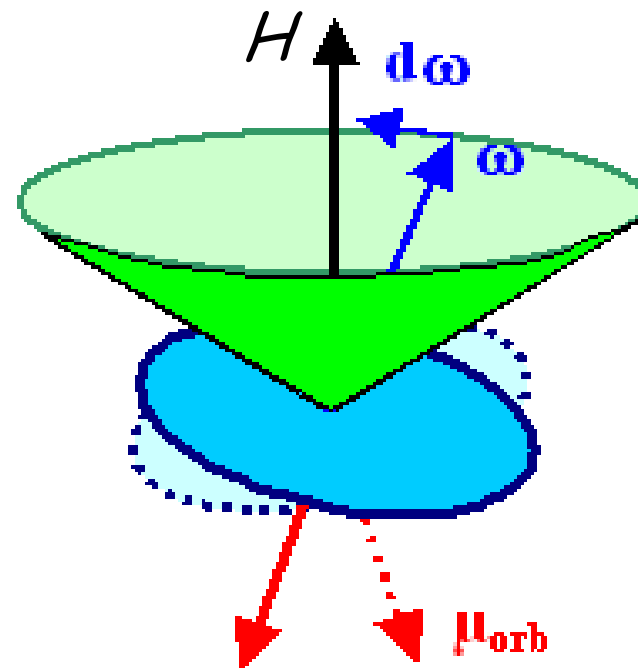
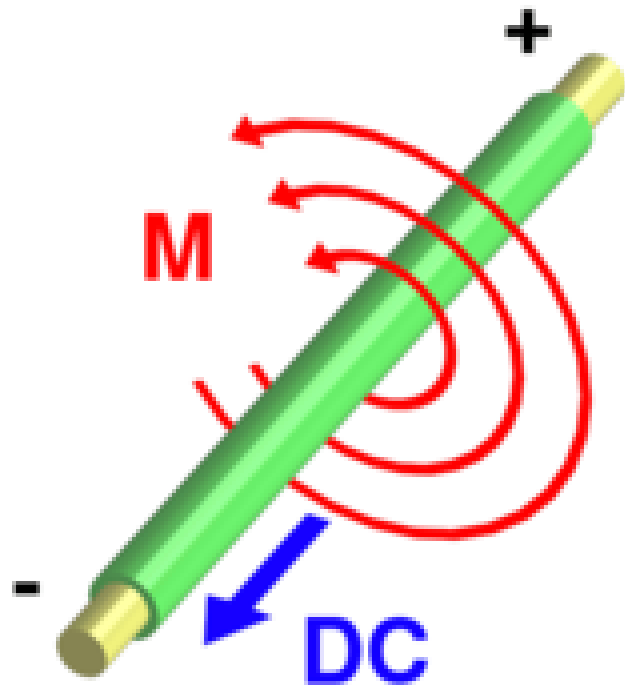


- Paramagnetism and diamagnetism can be distinguished experimentally by an apparatus like that above.
 - (a) no magnetic field
 - (b) paramagnetic substance appears to weigh more.
 - (c) diamagnetic substance appears to weigh less.

Source: INTERNET



Electric Current and Magnetic Field



Source: INTERNET





Magnetic Properties of Substances

$$\vec{\mathbf{B}} = \vec{\mathbf{B}}_0 + \vec{\mathbf{B}}'$$

$$\vec{\mathbf{B}}' = \chi \vec{\mathbf{B}}_0$$

$\vec{\mathbf{B}}$ is a resultant magnetic field

$\vec{\mathbf{B}}_0$ is an external magnetic field

$\vec{\mathbf{B}}'$ is an internal magnetic field

χ is a magnetic susceptibility





Dia- Para- and Ferro- Magnetic Substances

Diamagnetic substance - $\chi < 0$
no unpaired electrons in diamagnetic atoms

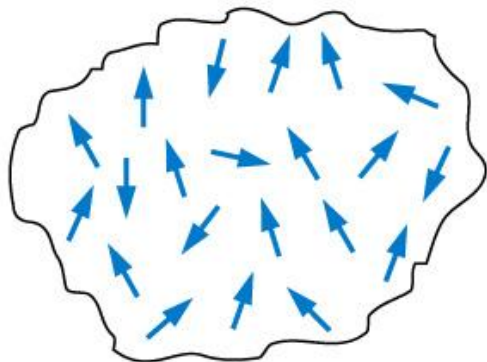
Paramagnetic substance - $\chi > 0$
unpaired electrons are present
in paramagnetic atoms

Ferromagnetic substance - $\chi \gg 1$
unpaired electrons are present in atoms, and
magnetic domains are present in this substance

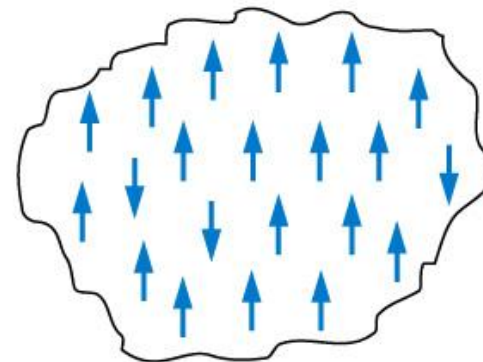


Para- and Ferro- Magnetic Substances

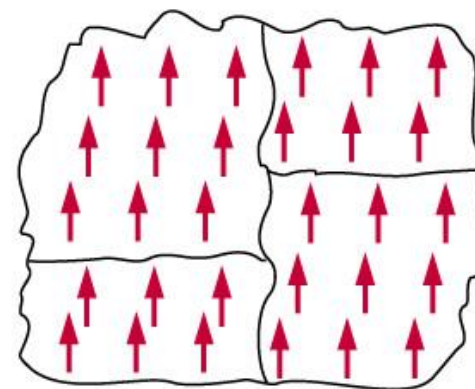
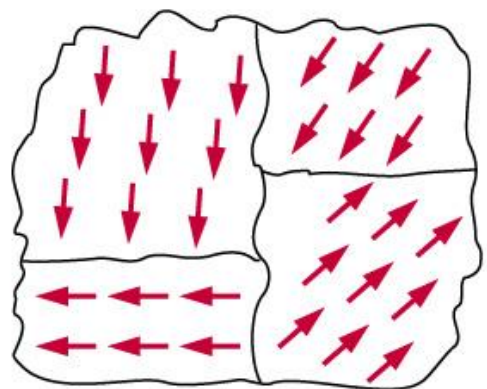
Magnetic field absent



In presence of magnetic field



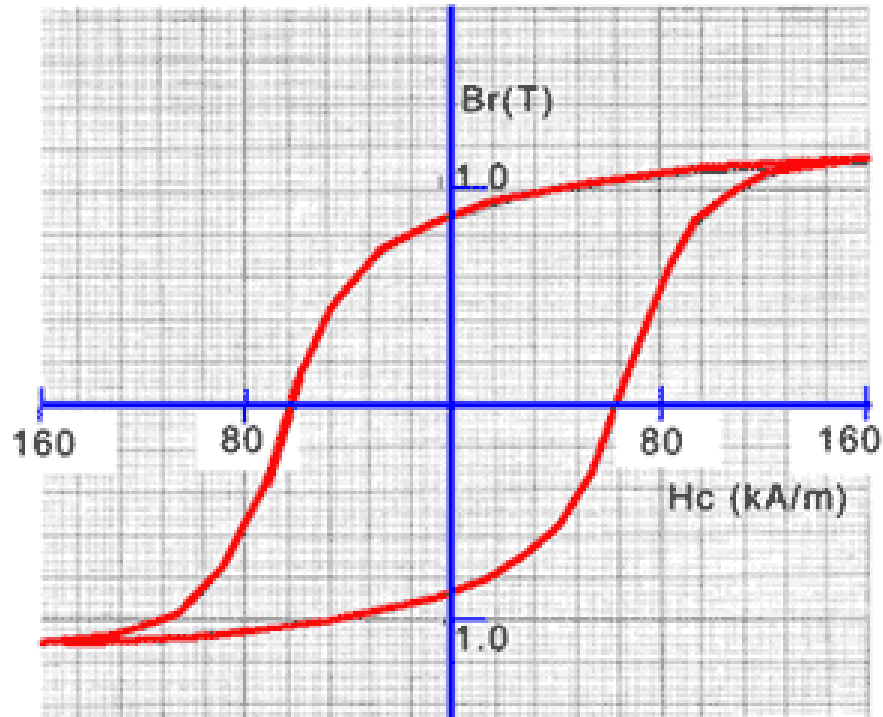
Paramagnetism



Ferromagnetism

Source: INTERNET

Magnets



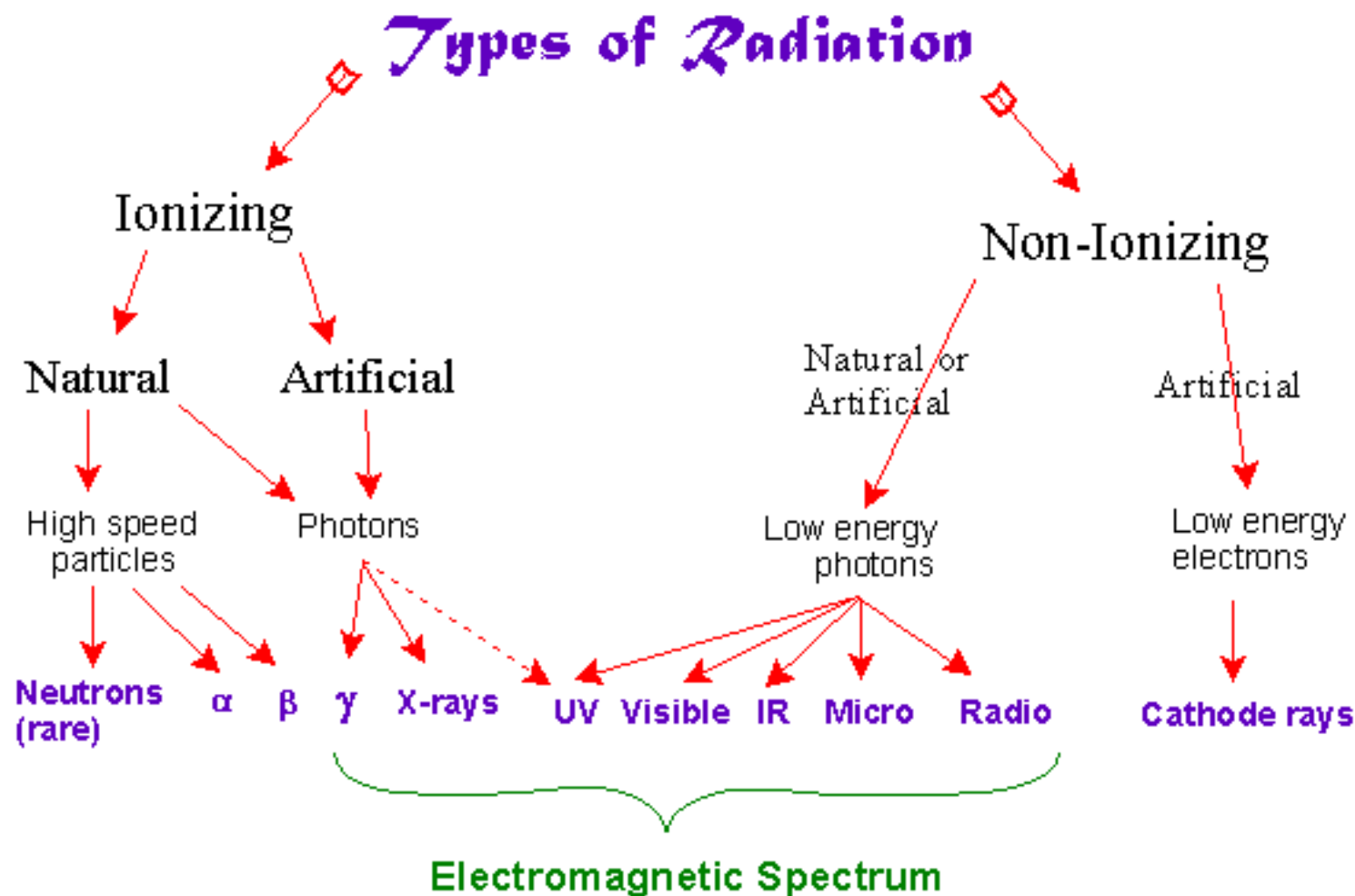
Permanent magnets



Electro-magnets

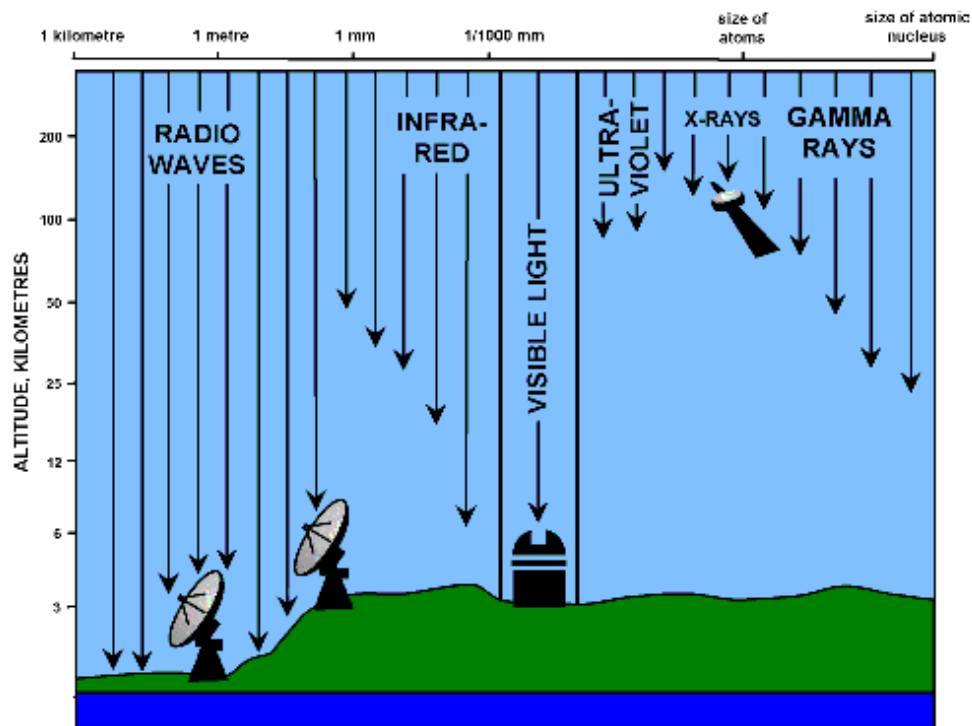
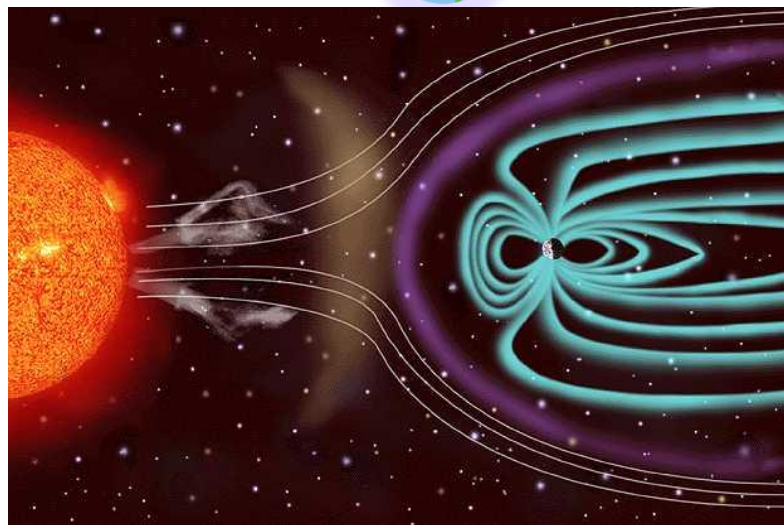
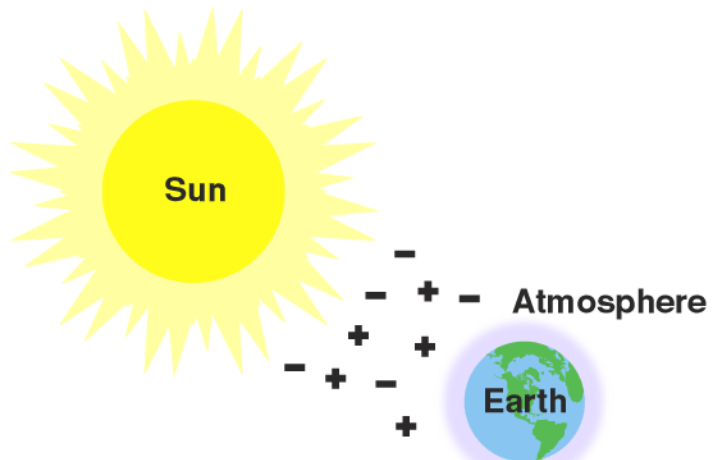
Source: INTERNET

Radiation



Cosmic Radiation – a General Concept

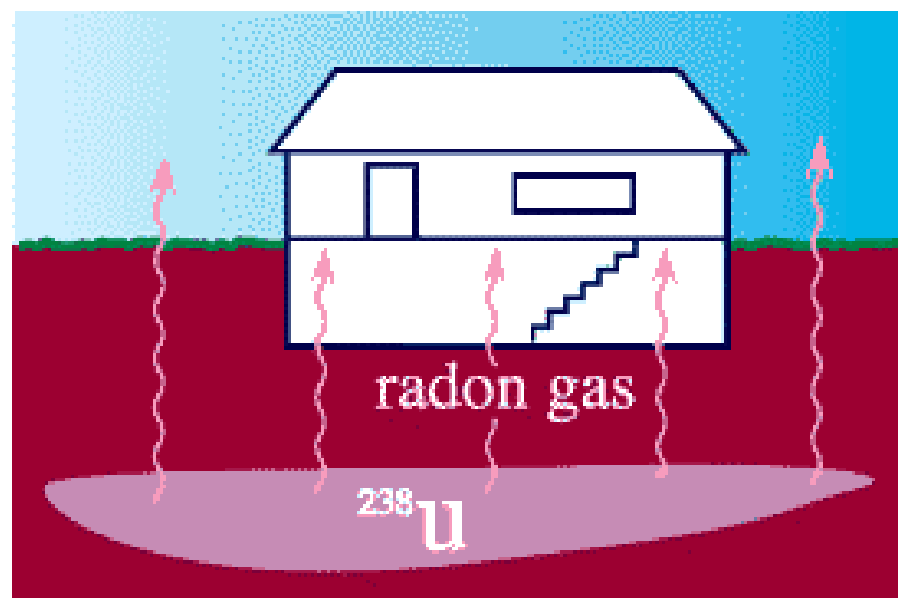
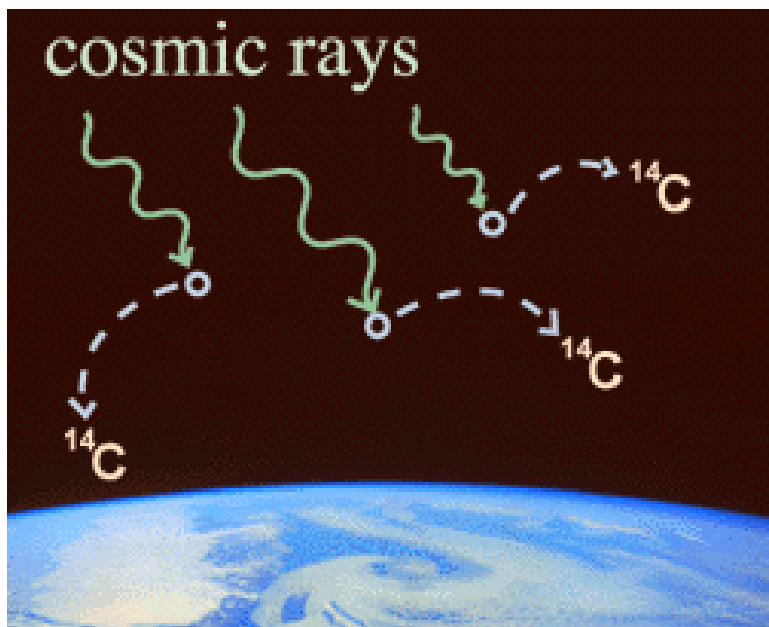
Cosmic Radiation



Source: INTERNET



Radiation of Background




Source: INTERNET

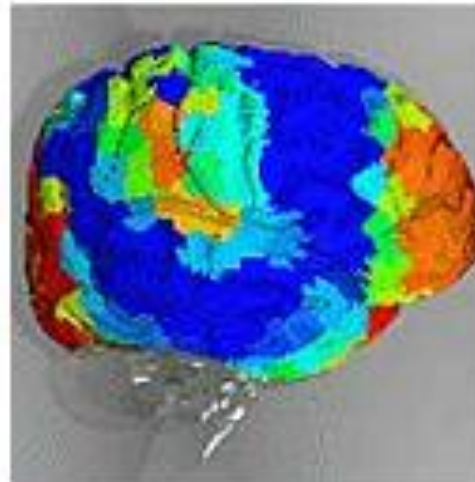




Nuclear Radiation from Human Activities



 A nuclear explosion



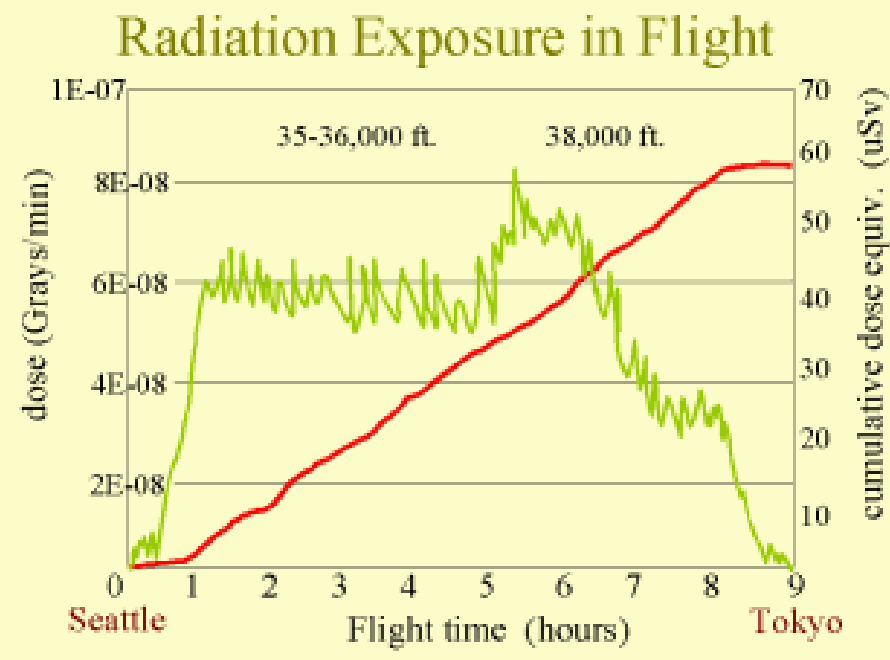
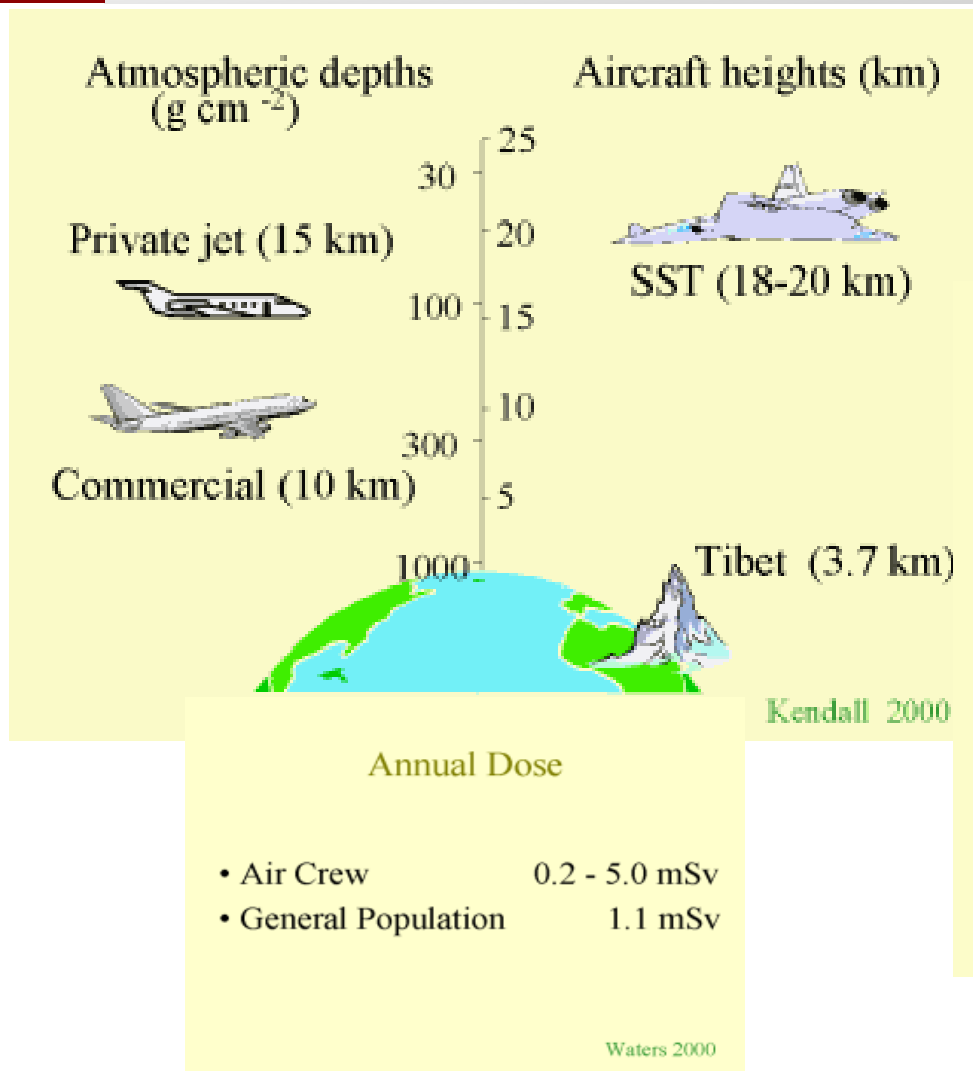
 Radiation medicine image of the brain.



 X-Rays are invisible radiation

Source: INTERNET

Is This Exposure Dangerous?



Source: INTERNET



Radiation Units

◆ Activity:

1 becquerel (Bq) = 1 particle emitted/sec

This is a very small unit (wouldn't do you much harm)

1 gram of radium has an activity of $3.7 \text{ E}10 \text{ Bq}$ or $3.7 \text{ E}4 \text{ MBq}$.

The old unit of radiation is Curie.

1 Curie = $3.7 \text{ E}10$ emissions/sec.

◆ Dosage:

1 gray (Gy) = 1 J/kg.

This means that we need to take into account the energy that the radiation has. Radiation with higher energy will cause more damage to the body.

Old unit = rad. 1 rad = $1/100 \text{ Gy}$.





Radiation Units

◆ Dose Equivalent

Even working with doses of radiation, there are still different effects on living cells. It depends on what type of radiation it is as well as what tissue is being irradiated. We use a quantity called "dose equivalent" to specify how much damage is actually done.

1 Seivert = 1 gray * quality factor.

Quality Factors: X-rays with energy of 200 keV = 1
gamma = 1; beta = 1-2; slow neutrons = 2-3; fast neutrons = 5-10;
alpha = 5-15; heavy ions = 10-20

Old unit = rems. 1 rem = 1/100 Sv.

LD ₅₀	pig	190 - 310 rems
	dog	240 - 320 rems
	human	about 300 rems
	rat	790 - 820 rems





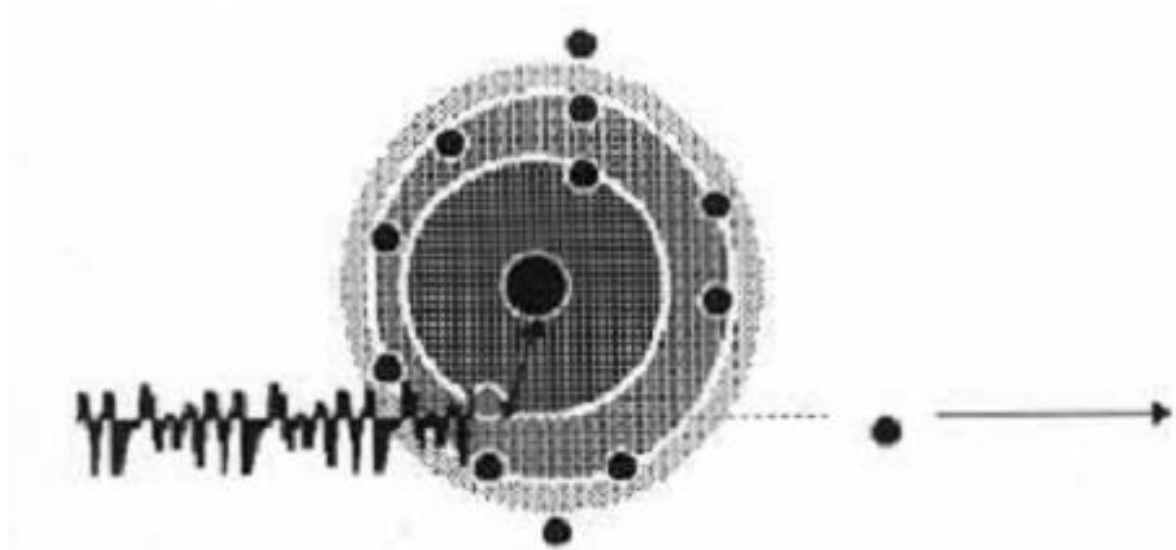
Radiation Stages

1. Physical stage (10^{-16} s) - ionisation of atoms and molecules, creation of secondary particles
2. Physio-chemical stage (10^{-13} s) - secondary reactions of ions and creation of stable molecules and unstable free radicals
3. Chemical stage (10^{-8} s) - reactions of free radicals
4. Biological stage (days, months years) - disrupted functions of destroyed biopolymers



Interaction of Gamma Radiation

Photoeffect

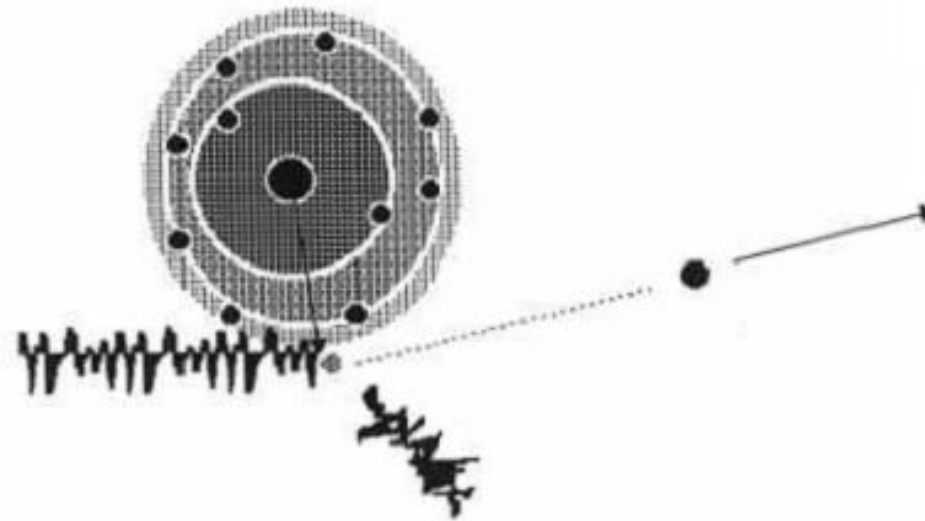


$$h\nu > W;$$

$$h\nu = W + E_k$$

Interaction of Gamma Radiation

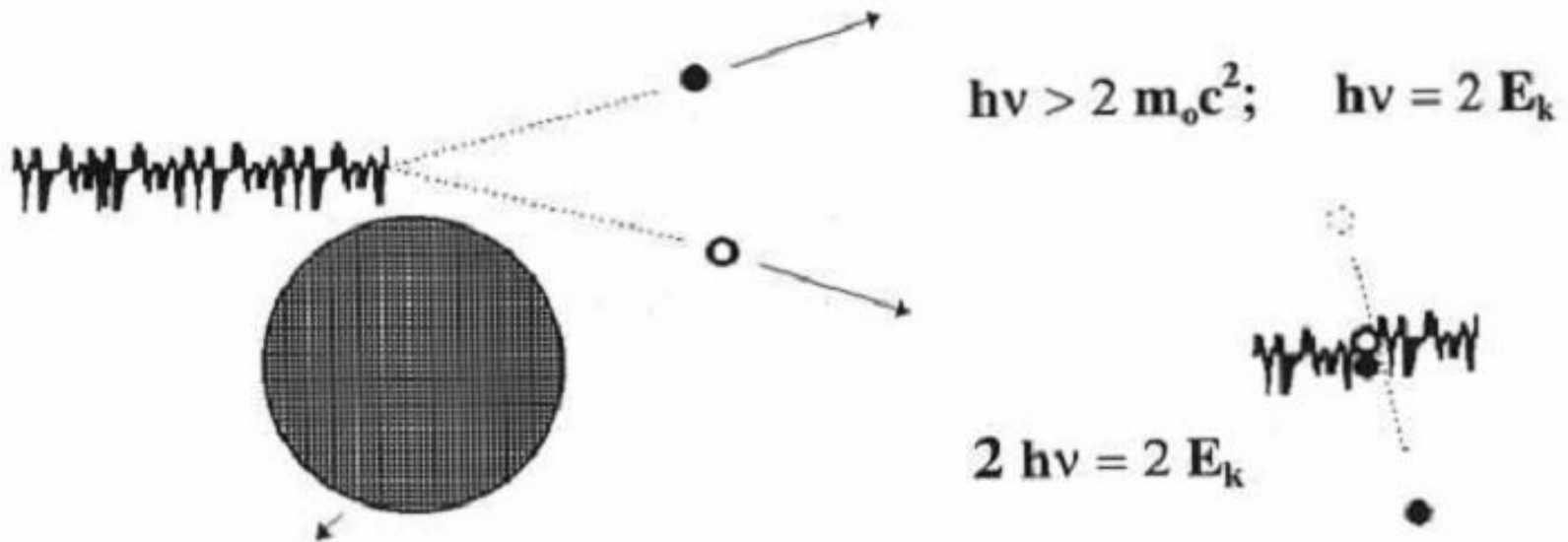
Compton effect



$$h\nu \gg W; \quad h\nu = h\nu' + E_k$$

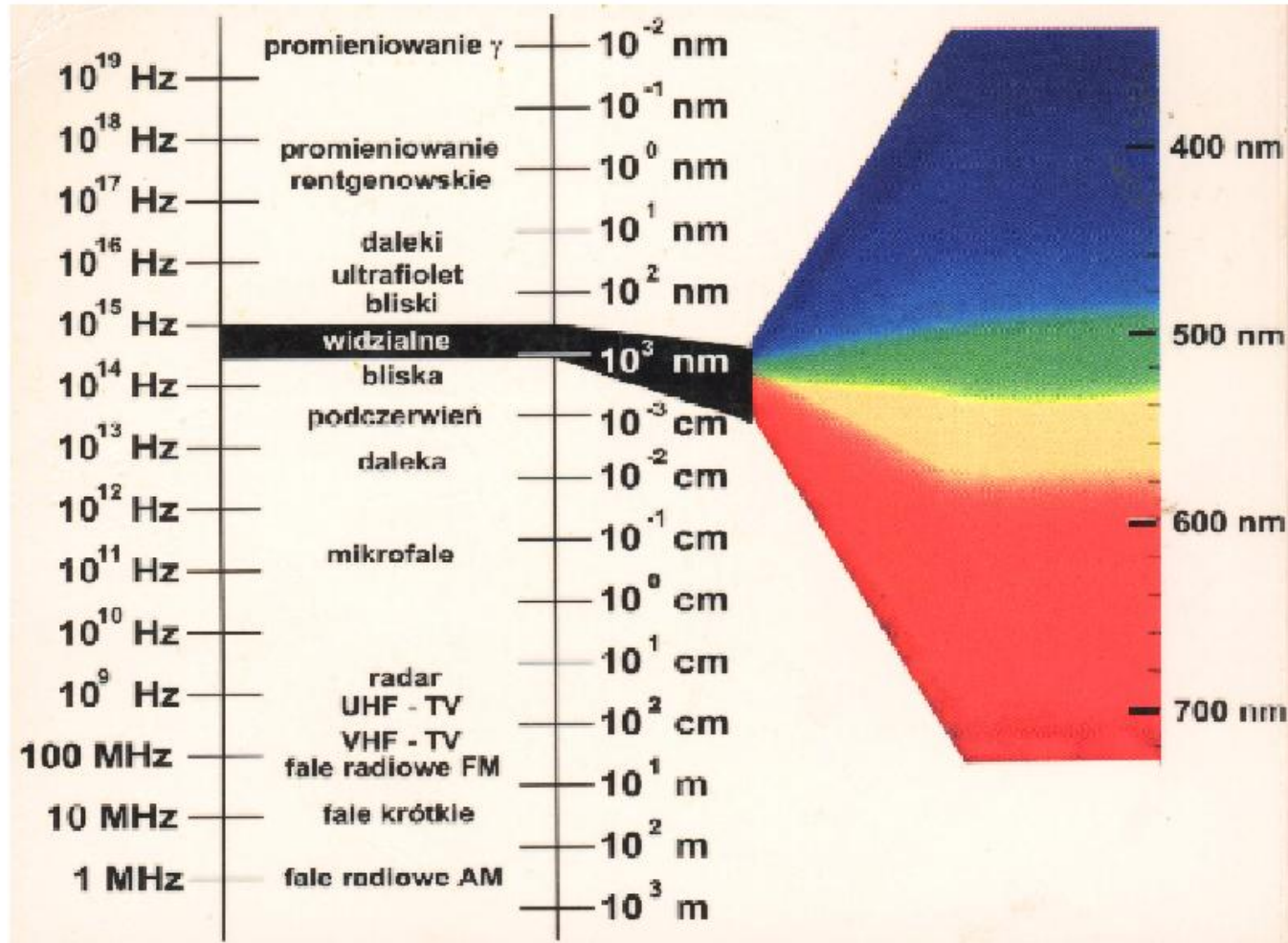
Interaction of Gamma Radiation

Pairs creation and annihilation





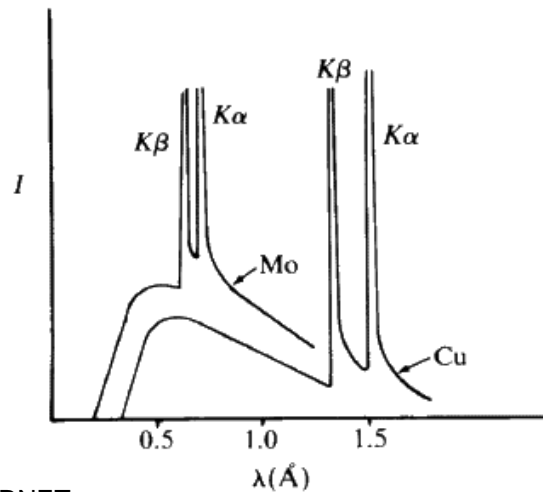
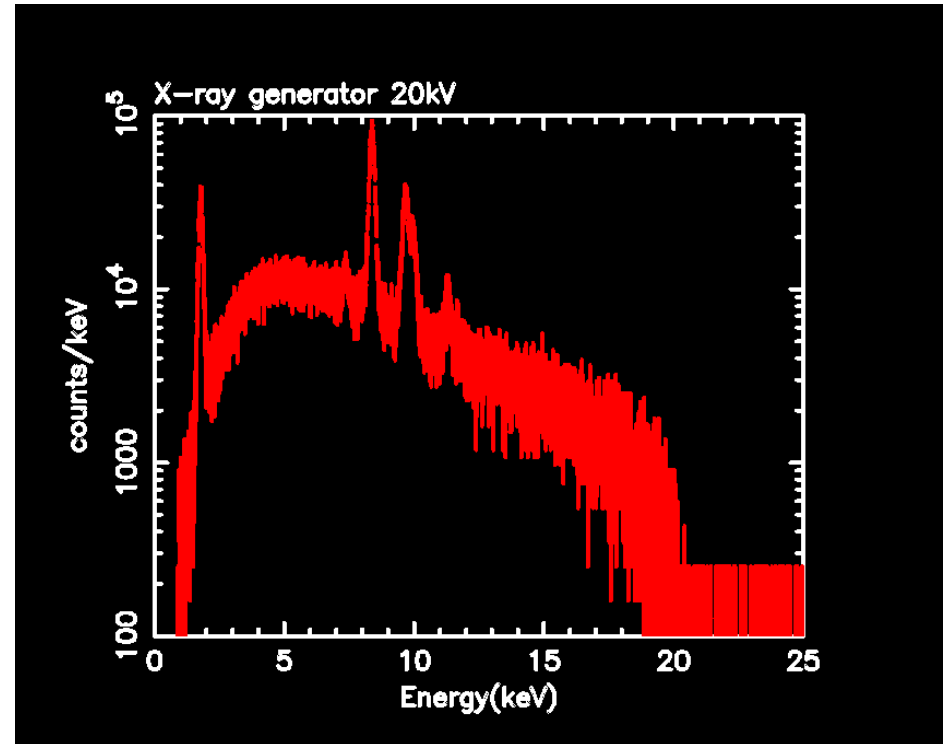
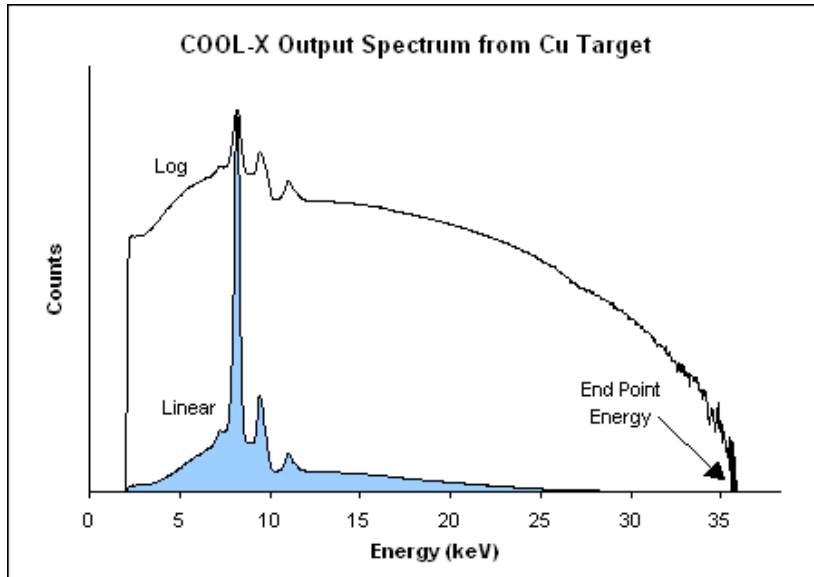
Spectra of Electromagnetic Radiation



Source: INTERNET



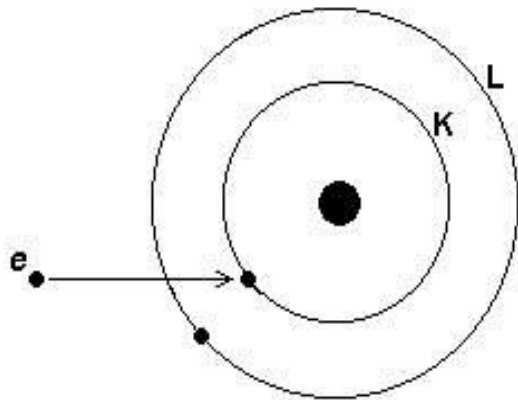
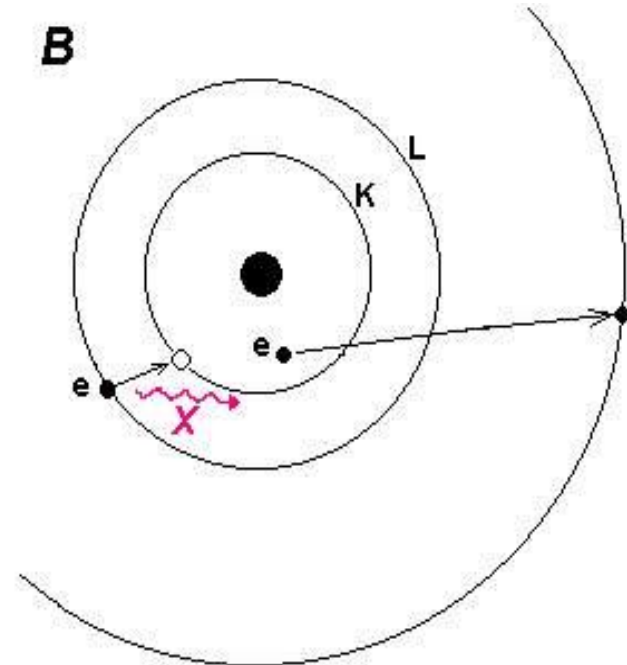
Spectra of X-rays



Source: INTERNET

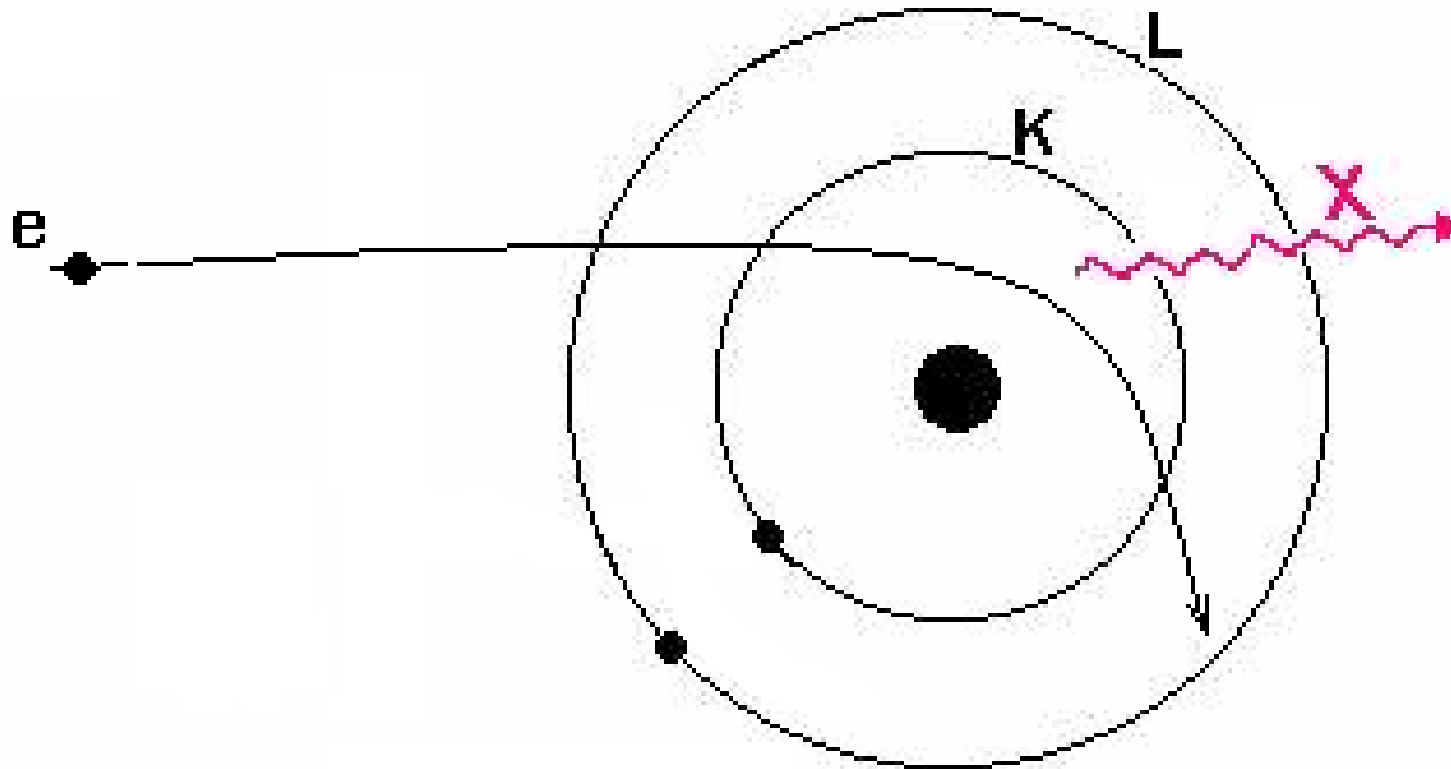
Characteristic X-rays

also K-shell emission or emission spectra

A**B**

Continuum Spectra of X-rays

also Bremsstrahlung or braking radiation





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