



KAPITAŁ LUDZKI
NARODOWA STRATEGIA SPÓJNOŚCI

UNIA EUROPEJSKA
EUROPEJSKI
FUNDUSZ SPOŁECZNY



BIOPHYSICS

**Prezentacja multimedialna współfinansowana przez
Unię Europejską w ramach
Europejskiego Funduszu Społecznego w projekcie pt.
*„Innowacyjna dydaktyka bez ograniczeń - zintegrowany
rozwój Politechniki Łódzkiej - zarządzanie Uczelnią,
nowoczesna oferta edukacyjna i wzmacniania zdolności
do zatrudniania osób niepełnosprawnych”***



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Lecture 5

POTENTIALS (5)

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Chemical Potential

Gibbs' formula in the differential form

$$dU = TdS + dW \quad (dU = dQ + dW; \quad dS = dQ/T)$$

various kind of work:

- volumetric work	$dW_v = - p dV$	
- mechanical work	$dW_m = F dl$	
- electric work	$dW_e = \varphi dq$	
- chemical work	$dW_c = \mu dn$	μ – chemical potential

generally the work is given by the formula:

$$dW = dW_v + dW_m + dW_e + dW_c = - p dV + F dl + \varphi dq + \mu dn$$

Gibbs' formula can be rewritten to:

$$dU = TdS - p dV + F dl + \varphi dq + \mu dn$$





Chemical Potential

Let assume: $dS = 0$; $dV = 0$; $dl = 0$; $dq = 0$

then: $dU = \mu dn$

or generally:

$$dU = \sum_{i=1}^m \mu_i dn_i$$





Chemical Potential

When system is doing the work ($-dW$) a chemical potential is diminishing ($-d\mu$)

$$-d\mu = pdV$$

The equation of state for ideal gas is given by the formula:

$$pV = nRT \quad R=8.31 \text{ J/mol deg} - \text{the universal constant}$$

for isothermal process ($dT = 0$) valid is the equality:

$$p dV = - V dp$$

then:

$$-d\mu = - V dp; \quad \text{and} \quad V = nRT/p$$

$$d\mu = nRT dp/p$$

$$\int d\mu = nRT \int dp/p$$

$$\mu = nRT \ln p \Big|_1^2 + W$$





Chemical Potential

$$\mu = nRT (\ln p_2 - \ln p_1) + \mu^\circ$$

$$\mu = + \mu^\circ + nRT \ln p_2 \quad \text{for } p_1 = 1$$

since $p = nRT/V$ and $n/V = c$

thus

$$\mu = + \mu^\circ + nRT \ln c$$





Electrochemical Potential

Electric charge of n mols of ions is described by the equation:

$$q = z n F$$

and change in the charge amount:

$$dq = z F dn$$

or for several types of ions:

$$dq = F \sum_{i=1}^m z_i dn_i$$

the Gibbs' formula will be expressed as:

$$dU = TdS - p dV + F dl + \varphi F \sum_{i=1}^m z_i dn_i + \sum_{i=1}^m \mu_i dn_i$$





Electrochemical Potential

Elements of the sum can be ordered:

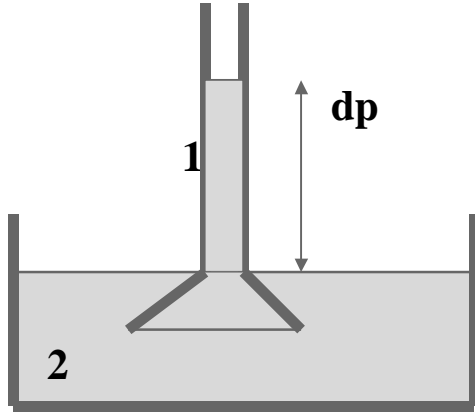
$$\phi F \sum_{i=1}^m z_i dn_i + \sum_{i=1}^m \mu_i dn_i = \sum_{i=1}^m dn_i (\mu_i - z_i F \phi)$$

$$\mu_i^e = \mu_i + z_i F \phi$$

$$\mu_i^e = \mu_i^{e0} + RT \ln c_i + z_i F \phi$$



Osmotic Pressure



$$\mu^{\circ}_1 + RT \ln c_1 = \mu^{\circ}_2 + RT \ln c_2$$

$$RT \ln c_1/c_2 = \mu^{\circ}_2 - \mu^{\circ}_1$$

$$\mu^{\circ}_2 - \mu^{\circ}_1 = V dp$$

subsystem 1: $\mu_1 = \mu^{\circ}_1 + RT \ln c_1$

subsystem 2: $\mu_2 = \mu^{\circ}_2 + RT \ln c_2$

$$\mu_1 = \mu_2$$

$$dp = (RT/V) \ln c_1/c_2$$

$$\Pi = (RT/V) \ln c_1/c_2$$



Osmotic Pressure

If the subsystem 2 contains pure solvent then $c_2 = 1$, and:

$$\Pi = (RT/V) \ln c_1$$

for low concentrations of solutions true is dependence:

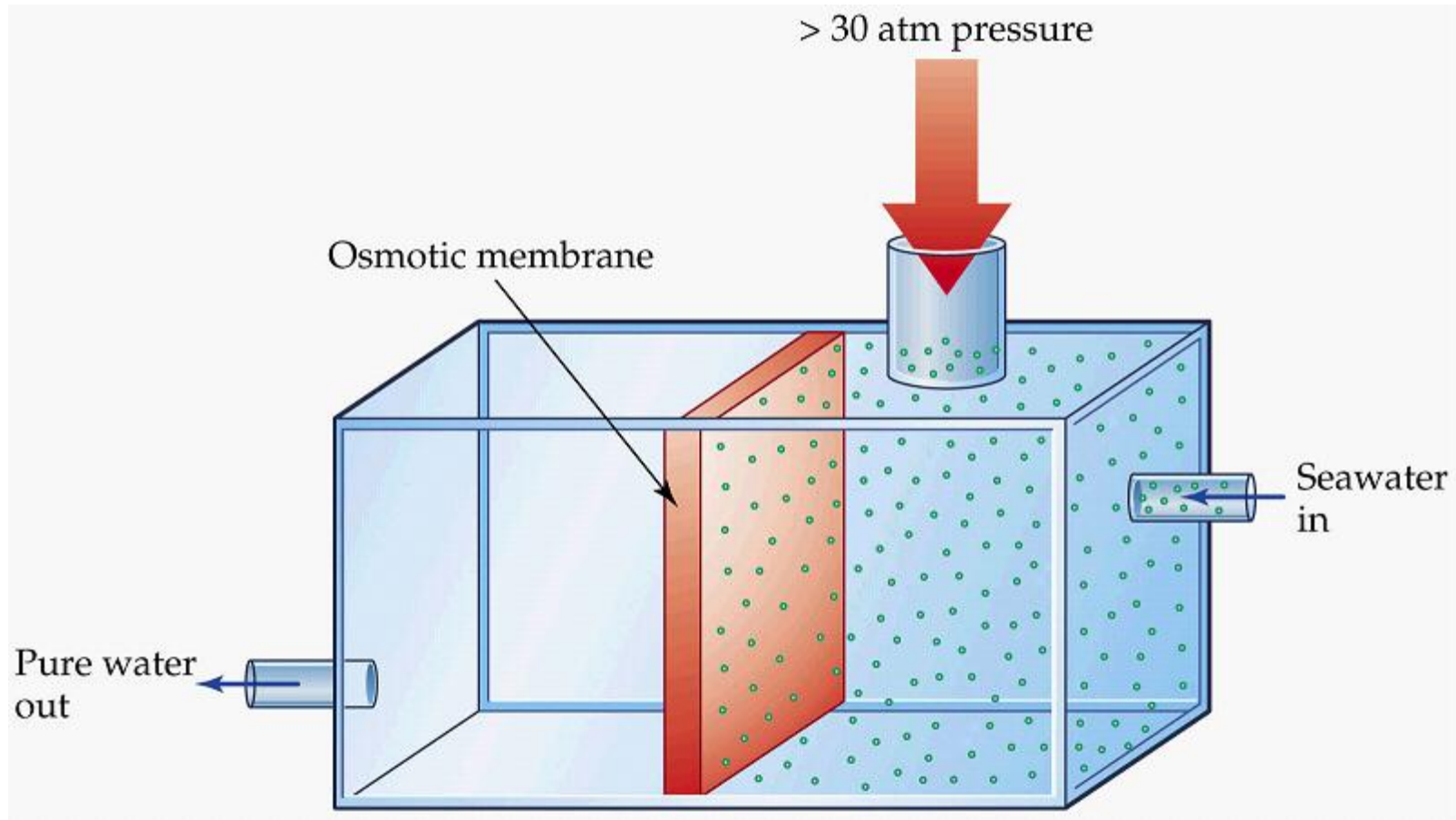
$$\Pi = RTc_1$$

Van't Hoff' formula





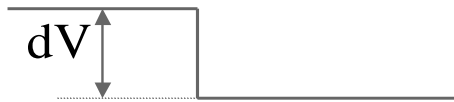
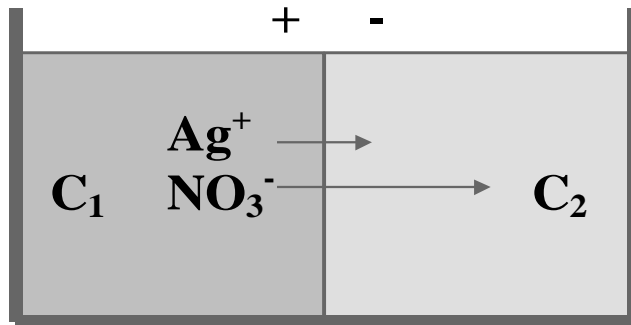
Osmotic Pressure



Source: INTERNET



Diffusive (Contact) Potential



For equilibrium exchange $J_e = 0$, and then:

$$dV = - d\mu L_{ed}/L_e$$

and assuming that $d\mu = 0$, one can write:

$$J_d = L_d d\mu + L_{de} dV$$

$$L_{ed}/L_e = J_d/J_e$$

$$J_e = L_{ed} d\mu + L_e dV$$

$$J_d = c (v^- - v^+) ; \quad J_e = z F c (v^- + v^+)$$



Diffusive (Contact) Potential

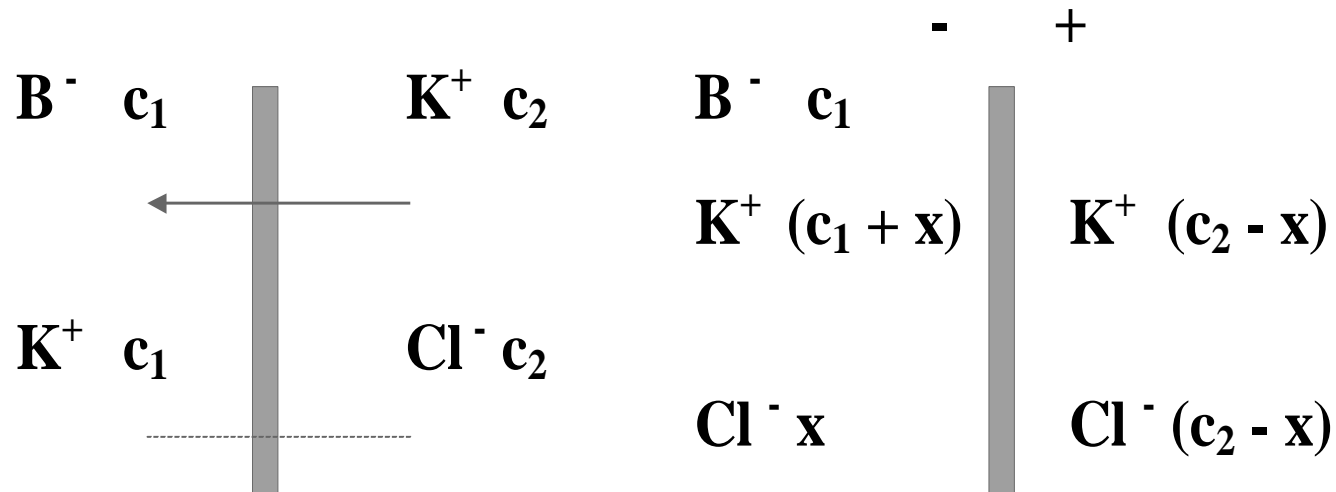
$$dV = \frac{v^+ - v^-}{v^+ + v^-} \frac{1}{z F} d\mu; \quad \mathbf{u = v/E - mobility}$$

$$dV = \frac{u^+ - u^-}{u^+ + u^-} \frac{RT}{z F} \ln \frac{c_1}{c_2}$$





Donnan's Equilibrium



Starting state

Equilibrium state

$$d\mu^e = 0$$





Donnan's Equilibrium

$$\mathbf{x} (\mathbf{c}_1 + \mathbf{x}) = (\mathbf{c}_2 - \mathbf{x}) (\mathbf{c}_2 - \mathbf{x})$$

$$\mathbf{x} / (\mathbf{c}_2 - \mathbf{x}) = (\mathbf{c}_2 - \mathbf{x}) / (\mathbf{c}_1 + \mathbf{x})$$

$$[\mathbf{Cl}_1^-] / [\mathbf{Cl}_2^-] = [\mathbf{K}_2^+] / [\mathbf{K}_1^+]$$

$$dV = (RT/F) \ln ([\mathbf{Cl}_1^-] / [\mathbf{Cl}_2^-]) = (RT/F) \ln ([\mathbf{K}_2^+] / [\mathbf{K}_1^+])$$





Membrane Potential

INTERIOR	EXTERIOR	c_I / c_E	E (mV)
Na⁺ 9.2 mM	Na⁺ 120 mM	13:1	+ 67
K⁺ 140 mM	K⁺ 2.5 mM	1:56	- 102
Cl⁻ 4.0 mM	Cl⁻ 120 mM	30:1	- 86

Nerst's equation (membrane potential)

$$dV = \frac{RT}{zF} \ln \frac{c_z}{c_w}$$





Membrane Potential

Goldman's equation (with permeability coefficients)

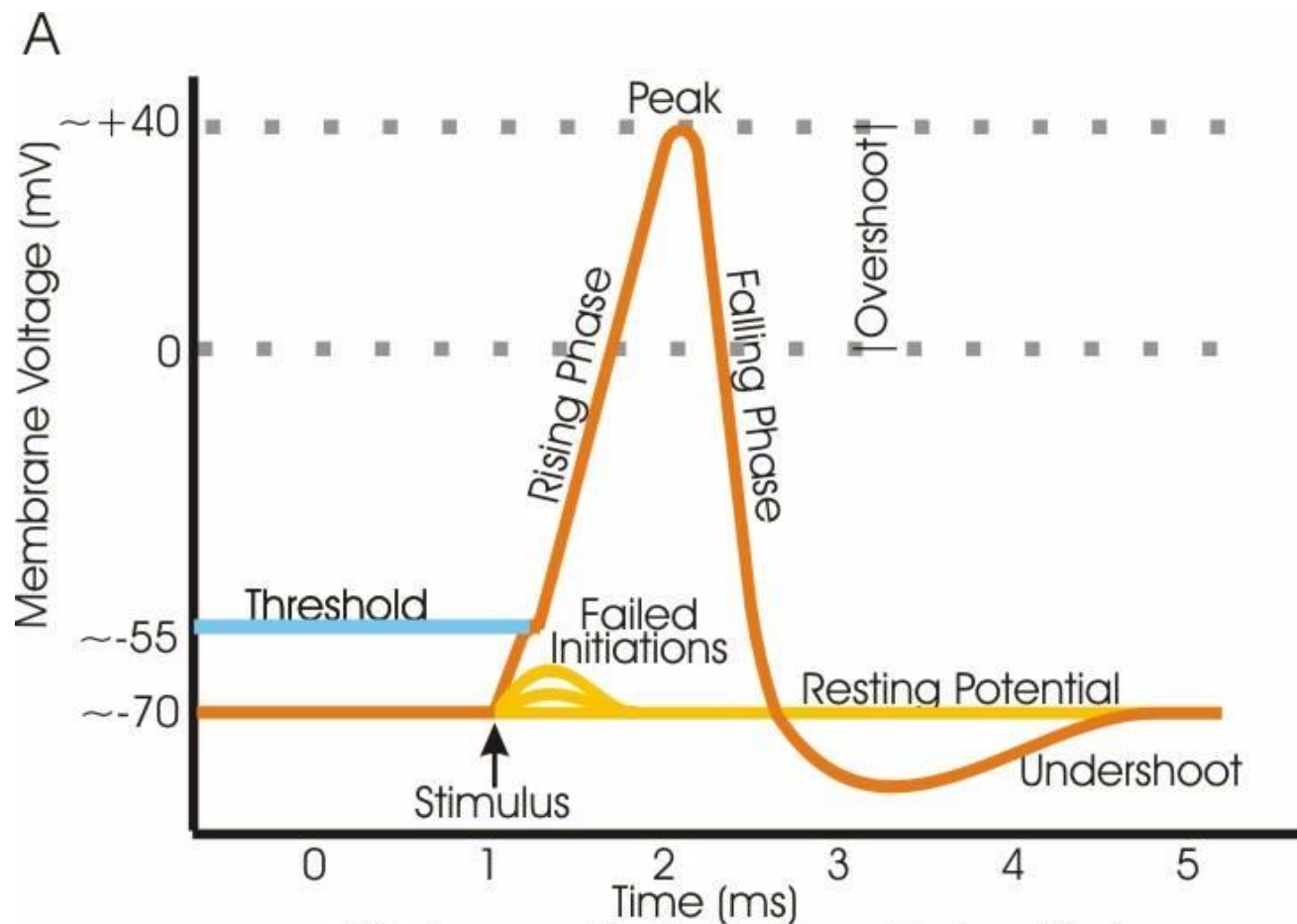
$$dV = \frac{RT}{zF} \ln \frac{P_K [K_I^+] + P_{Na} [Na_I^+] + P_{Cl} [Cl_E^-]}{P_K [K_E^+] + P_{Na} [Na_E^+] + P_{Cl} [Cl_I^-]}$$

For permeability coefficients $P_K : P_{Na} : P_{Cl} = 1 : 0.04 : 0.05$, and concentrations as on previous transparency, the calculated membrane potential is about **-90 mV**





Action Potential



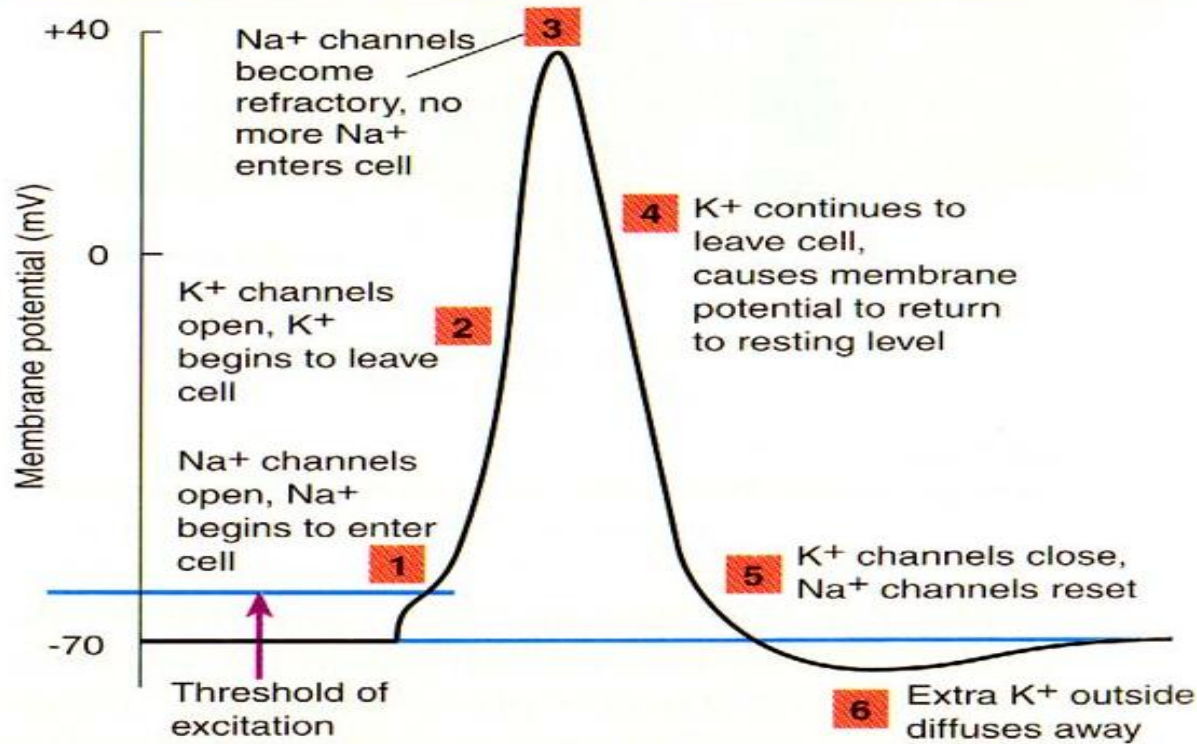
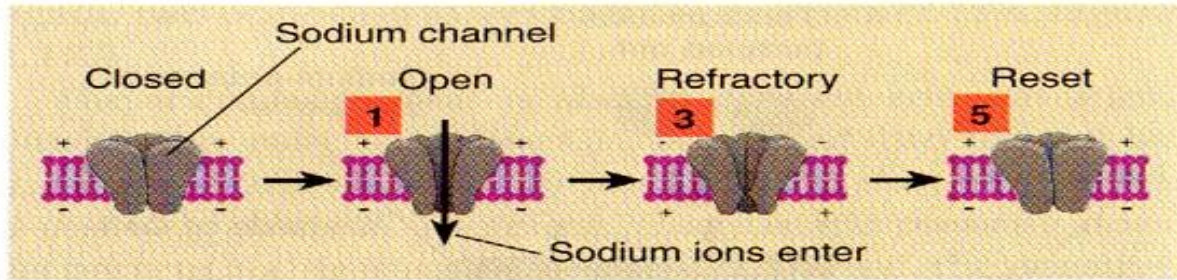
“Schematic” Action Potential

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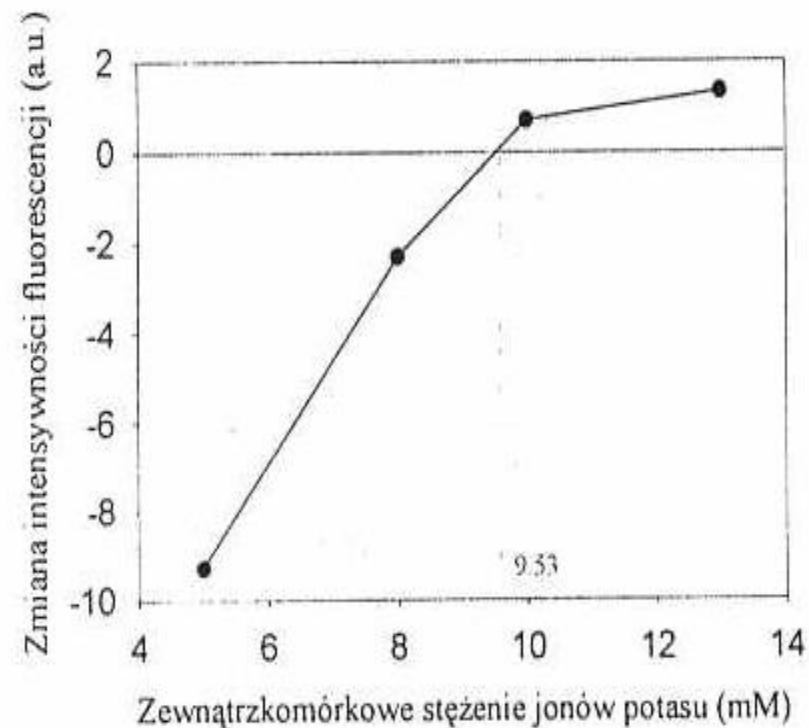
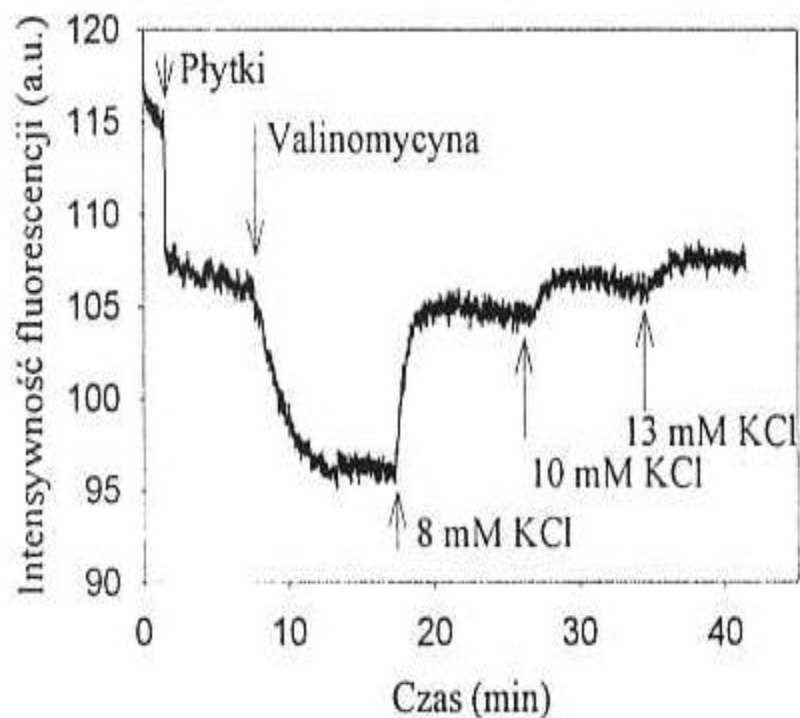
Action Potential



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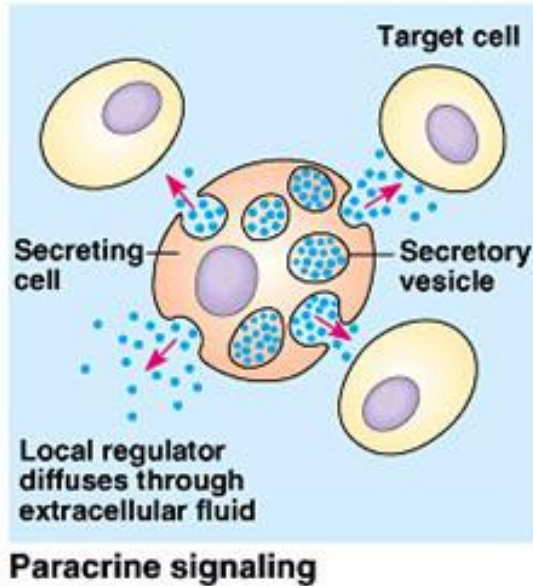


Measure of Membrane Potential

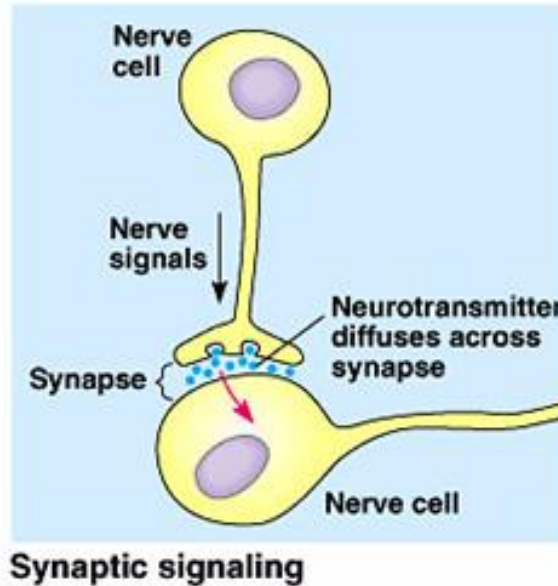




Cellular Signalling



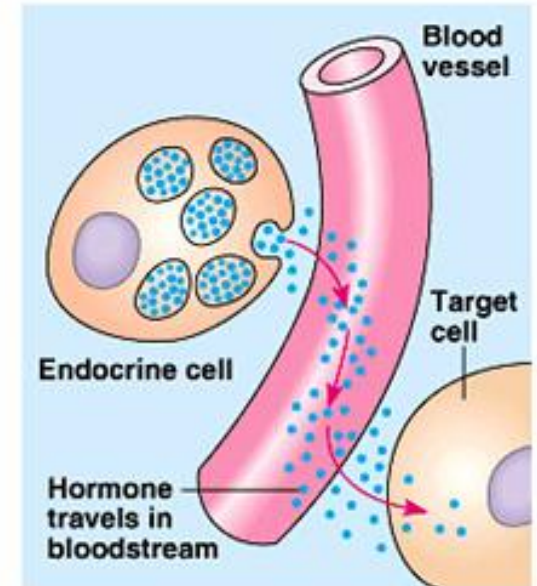
Paracrine signaling



Synaptic signaling

paracrine signaling

(a) Local signaling



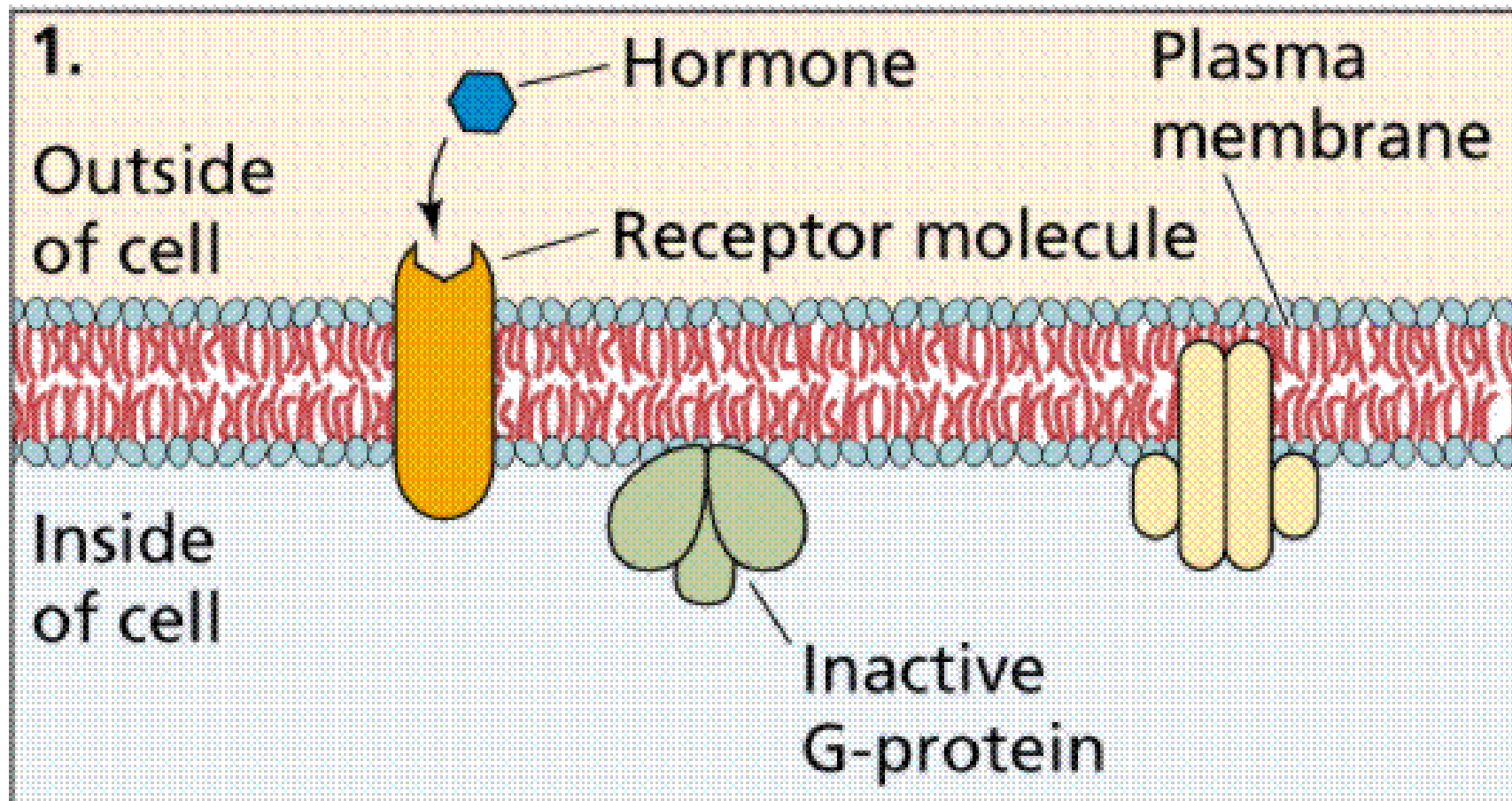
(b) Long distance (hormonal) signaling

endocrine signaling

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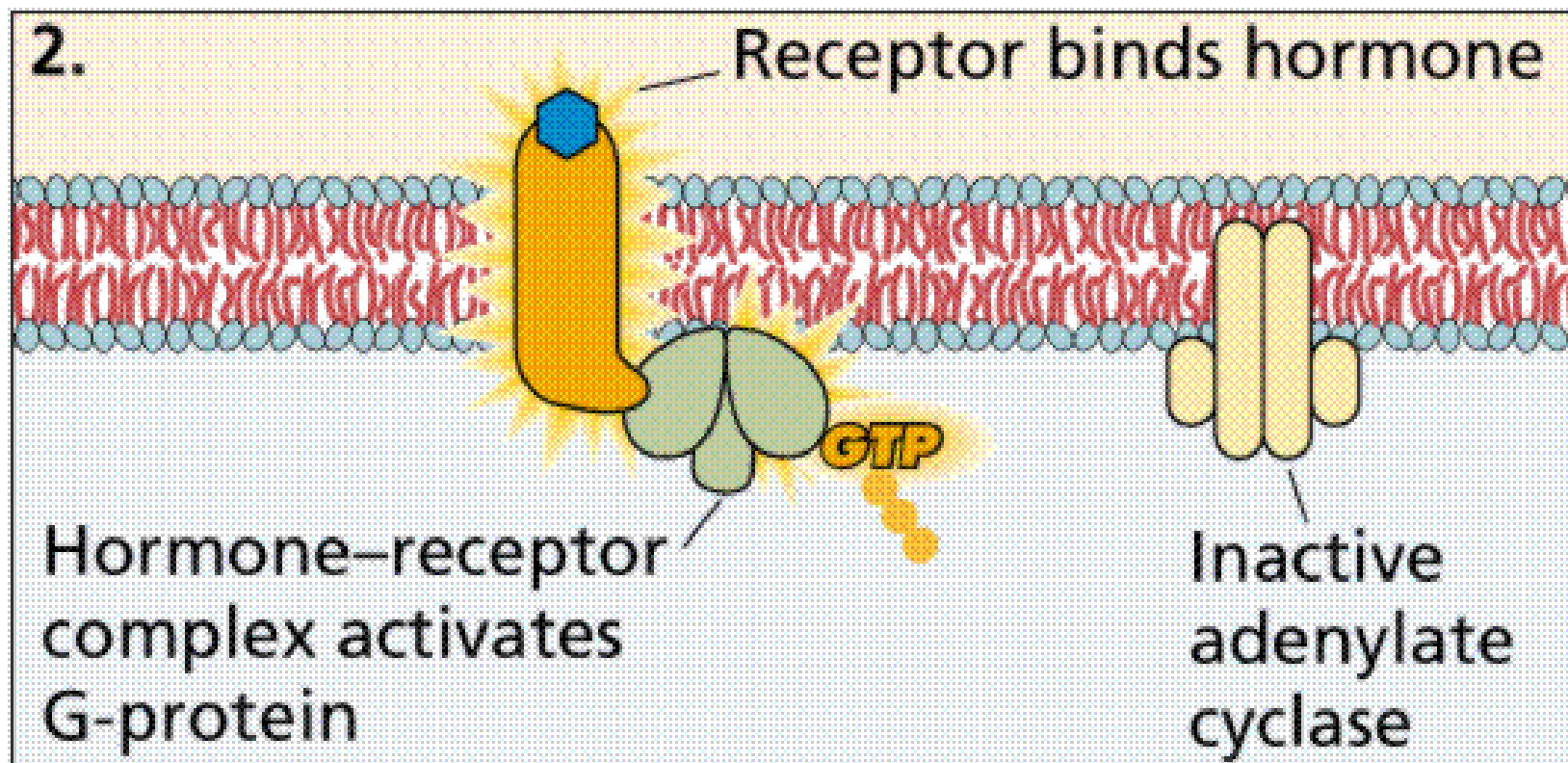


Transmembrane Signalling



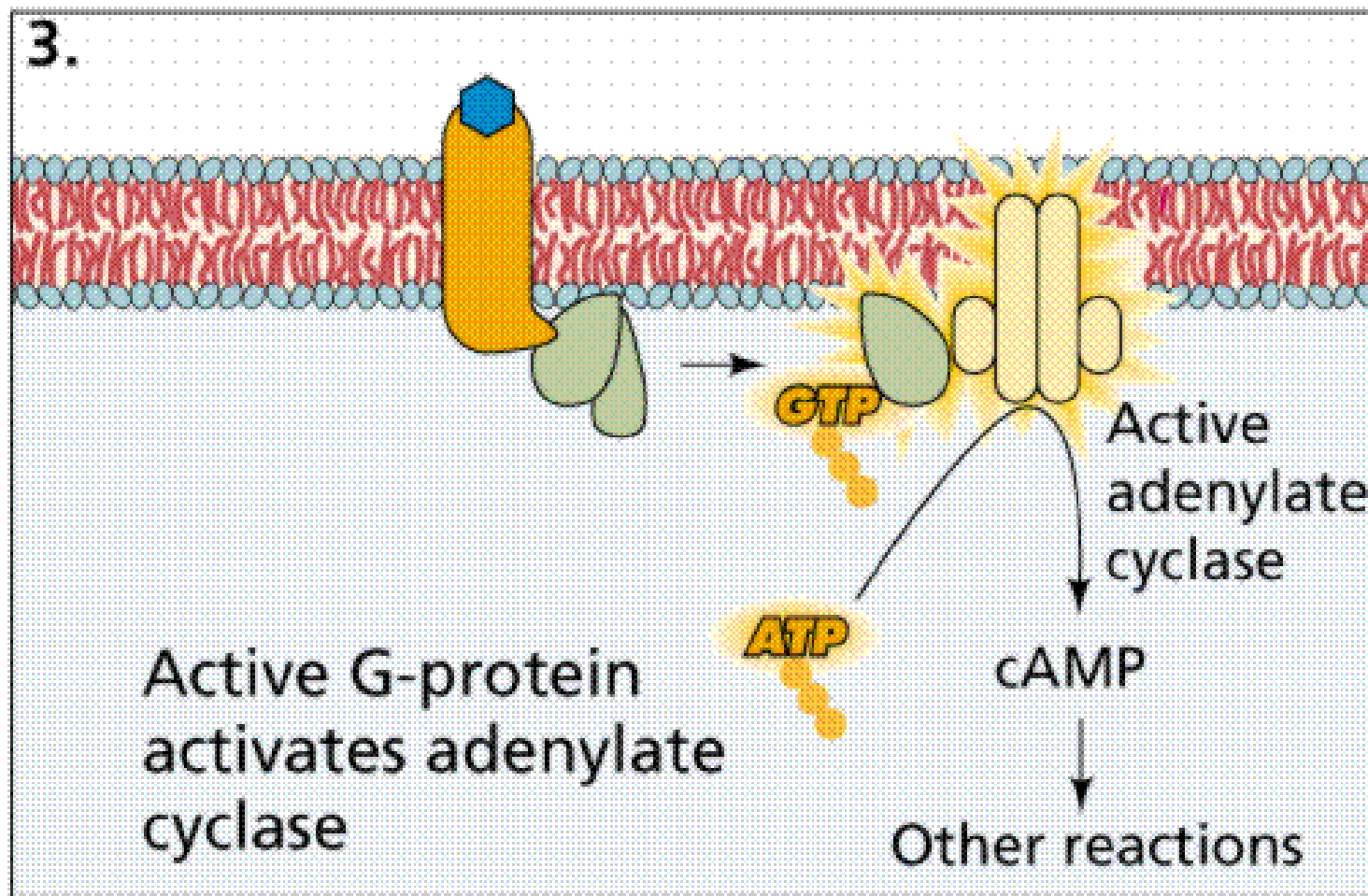
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Transmembrane Signalling



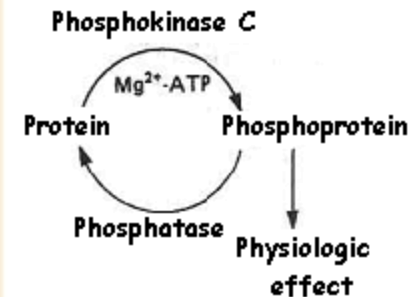
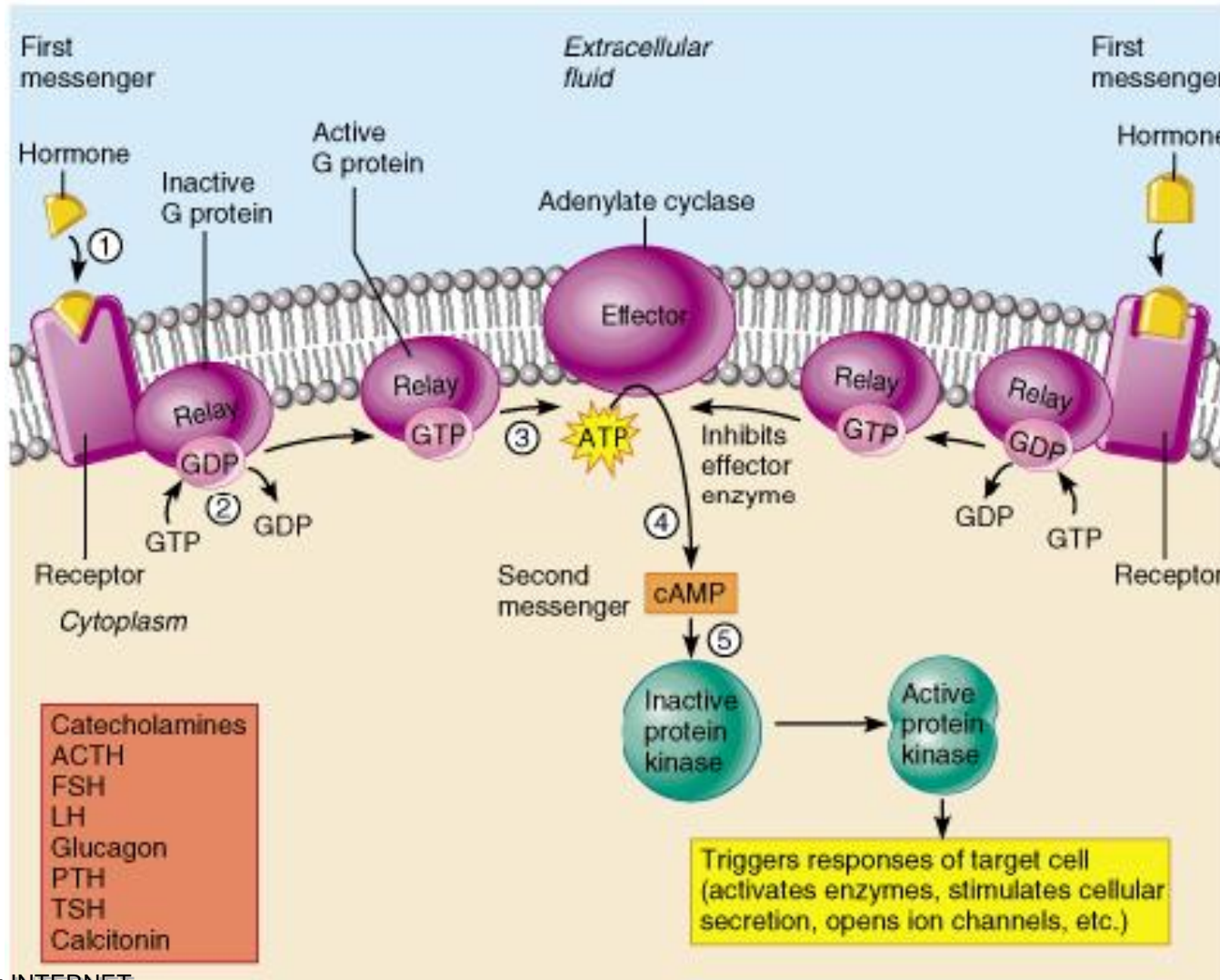
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Intracellular Signalling



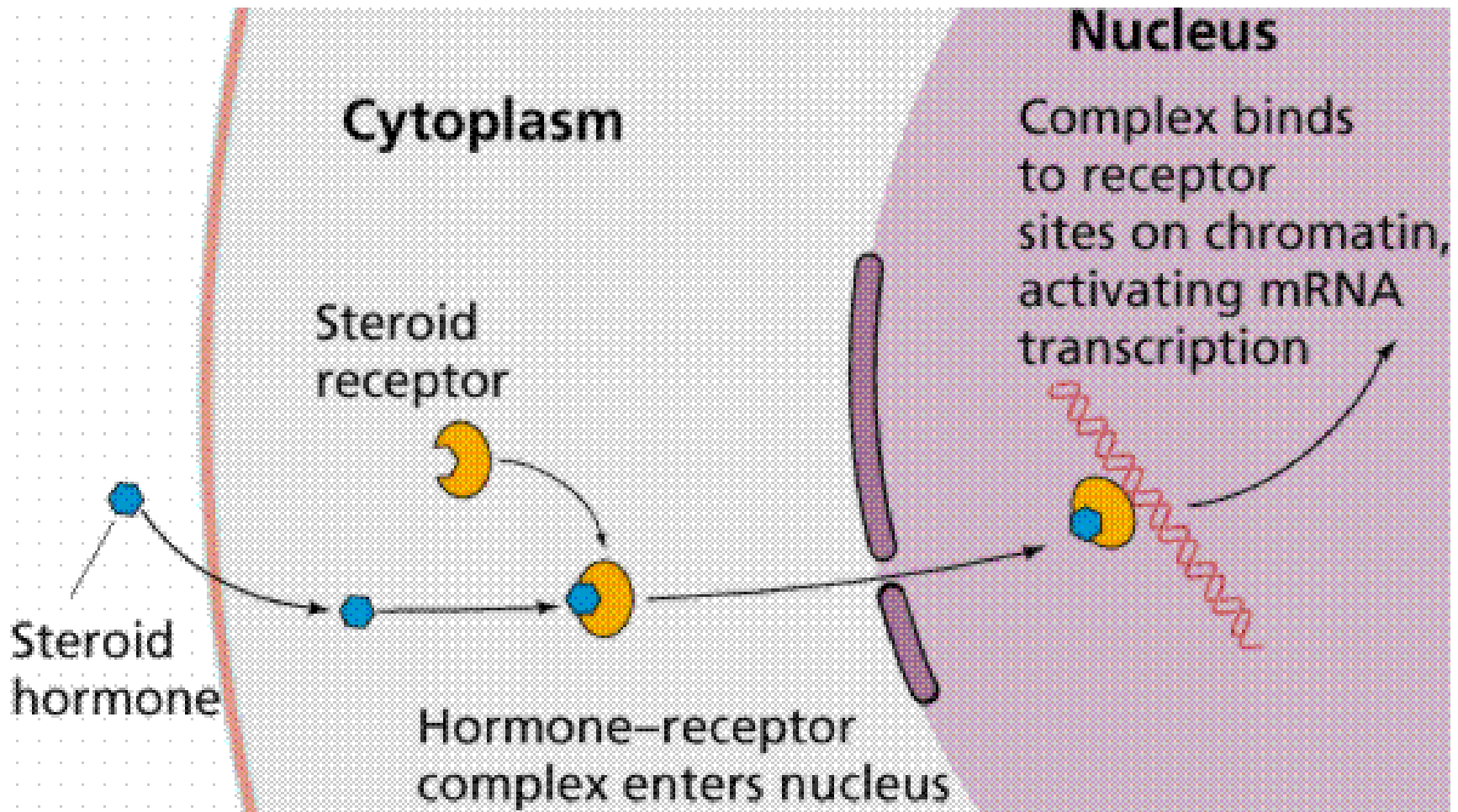
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Intracellular Signalling



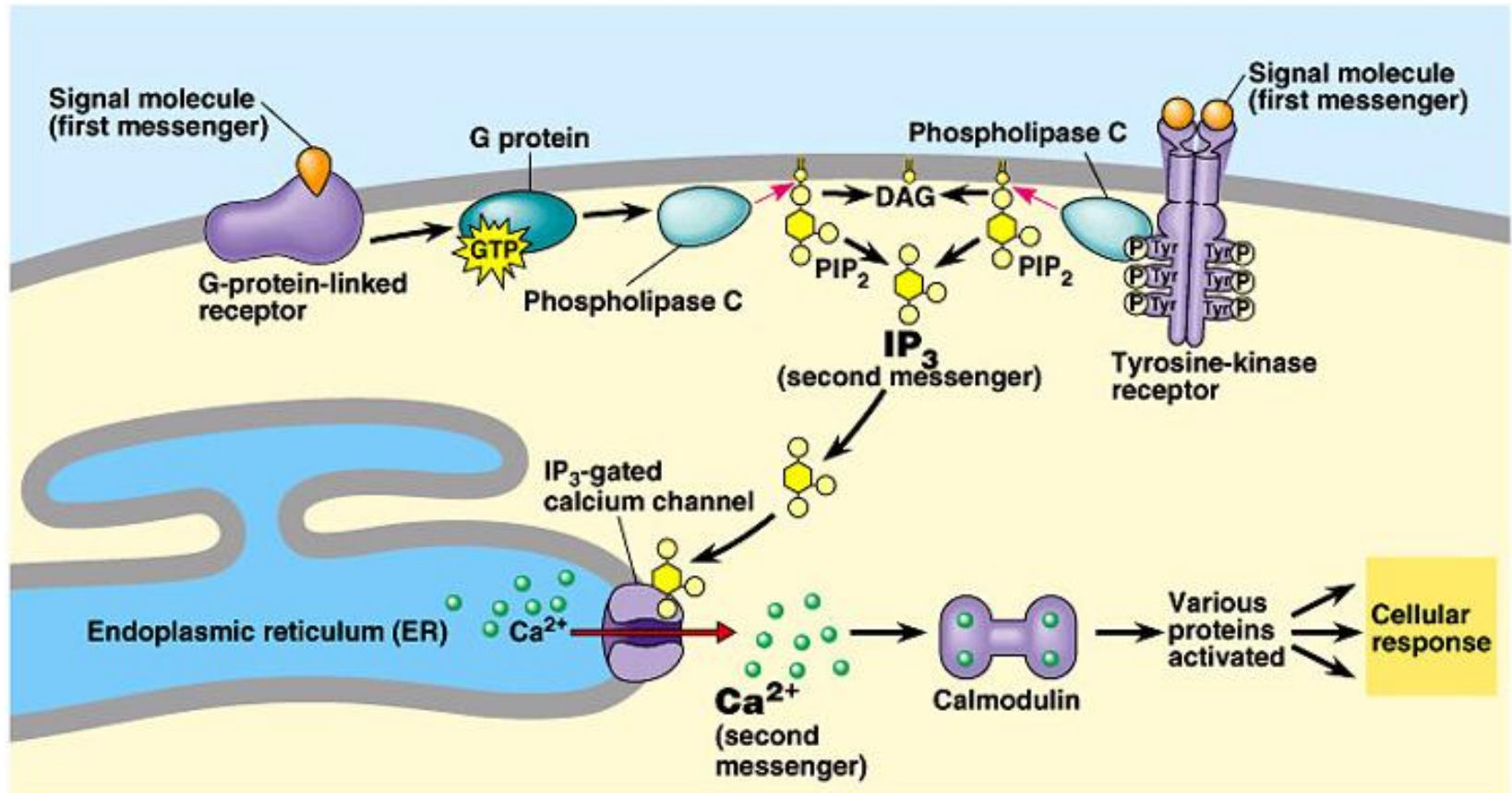
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Intracellular Signalling



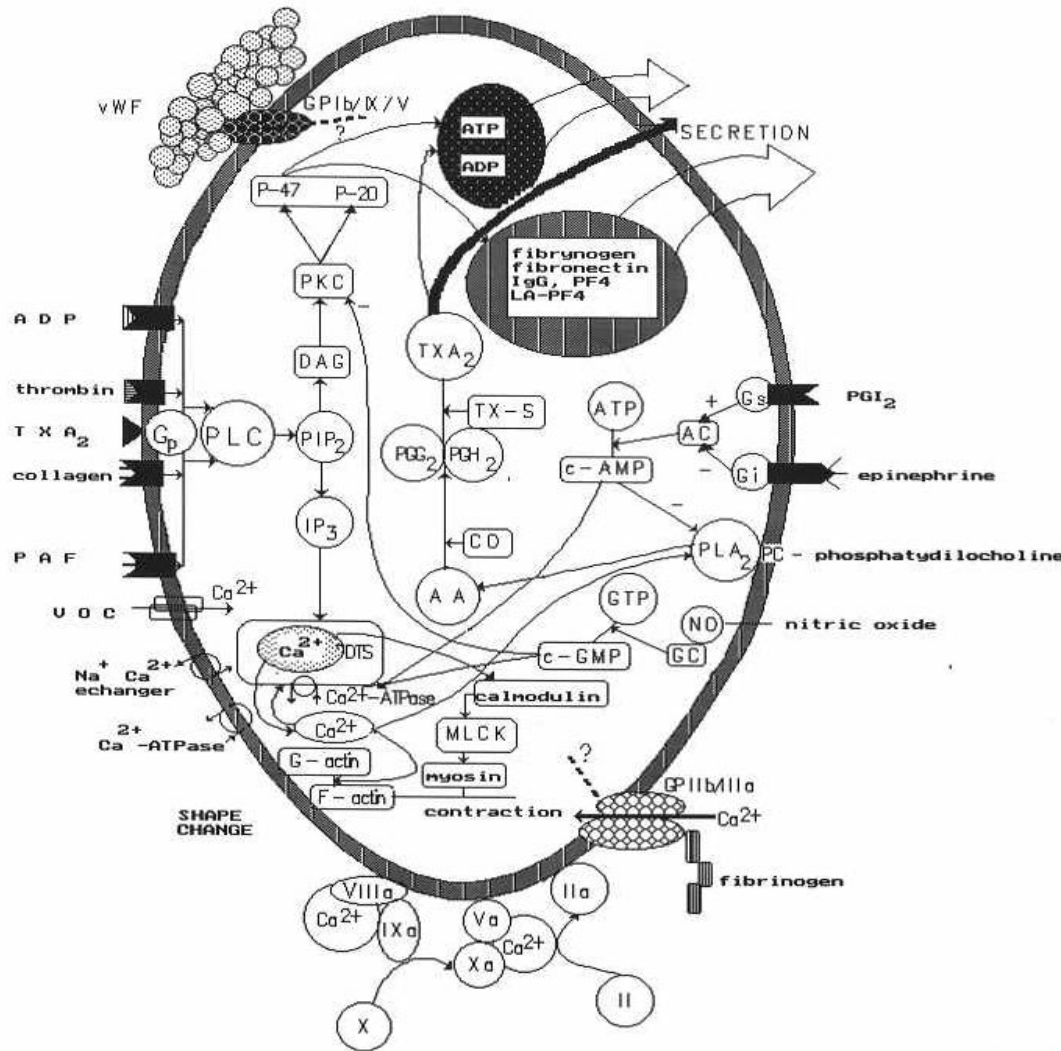
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Calcium Signalling



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Intracellular Signalling



- AC - adenylyate cyclase
- DAG - diacylglycerol
- PKC - protein kinase C
- PG - prostaglandins
- PLC - phospholipase C
- AA - arachidonic acid
- CO - cyclooxygenase
- TXA₂ - thromboxane A₂
- GC - guanylate cyclase
- PIP₂ - phosphatidylinositolbiphosphate
- IP₃ - inositoltriphosphate
- DTS - dense tubular system
- PLA₂ - phospholipase A₂
- TX-S - thromboxane synthetase
- MLCK - myosin light-chain kinase
- G_p, G_s, G_i - GTP binding proteins



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